

PHYSICAL: Kinetic Molecular Theory

Learning Objectives

- At standard atmospheric conditions, all gases have approximately the same kinetic energy, and are made up of molecules that behave as independent particles (the postulates of the kinetic molecular theory).
- At standard atmospheric conditions, all molecules have the same average kinetic energy.
- At standard atmospheric conditions, all molecules have the same average speed.

Model 1: Representation of Gases in Different Phases

The following figure illustrates the representation of particles in three different phases of matter.



Figure 1: Representation of particles in different phases of matter.

At standard atmospheric conditions, all gases have approximately the same kinetic energy, and are made up of molecules that behave as independent particles (the postulates of the kinetic molecular theory).

Key Concepts

1. All gases have the same average kinetic energy at standard atmospheric conditions, and are made up of molecules that behave as independent particles (the postulates of the kinetic molecular theory).
2. At standard atmospheric conditions, all molecules have the same average kinetic energy, and are made up of molecules that behave as independent particles (the postulates of the kinetic molecular theory).

Table 1: Representation of particles in different phases of matter.

	Gas	Liquid	Crystalline Solid
How the particles are arranged:	Particles are widely spaced and move randomly. They are not held together by any forces.	Particles are closely packed and move randomly. They are held together by weak forces.	Particles are closely packed and move randomly. They are held together by strong forces.
How the particles move:	Particles move rapidly in all directions. They are not held together by any forces.	Particles move rapidly in all directions. They are held together by weak forces.	Particles move rapidly in all directions. They are held together by strong forces.

Pogil Answer Kinetic Molecular Theory

Roman Wölfel



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