

This is like a distribution in math. $2(x+3) \rightarrow 2x+6$
 Everything inside of the parentheses is multiplied by two.

- b. What does the subscripted "2" outside the parentheses of the chemical formula tell you about the compound?

The subscript "2" outside of $\text{Ba}(\text{NO}_3)_2$ the parentheses indicate that there are two of the polyatomic ion nitrate (NO_3^-).

20. How many atoms of each element are in one formula unit of ammonium phosphate, $(\text{NH}_4)_3\text{PO}_4$?

nitrogen
3

hydrogen
12

phosphorus
1

oxygen
4

21. A student writes the chemical formula for the ionic compound calcium hydroxide as CaOH_2 .

- a. Write the chemical formula for each ion in the compound.

Calcium: Ca^{+2}

Hydroxide: OH^-

- b. Why is the student's chemical formula for the compound calcium hydroxide wrong?

The subscript of "2" after hydrogen (H) indicate two hydrogens, not two hydroxide polyatomic ions. Parentheses should be placed around OH^- and a subscript of "2" should be placed outside of the parentheses.

22. Many of the chemical formulas in Model 3 include parentheses. Which one of the following rules summarizes the appropriate use of parentheses in ternary ionic compounds? For the three rules that do not apply in all cases, show at least one counter example from the chemical formulas in Model 3.

Parentheses are used around any ion that is used more than once in a formula unit.

↳ Not true - $\text{Al}_2(\text{CO}_3)_3$ - Al has more than one atom.

Parentheses are used around any polyatomic ion.

↳ Not true - KNO_3 - NO_3^- has no parentheses because there is only one.

Parentheses are used around any polyatomic ion used more than once in a formula unit.

Parentheses are only used around polyatomic anions used more than once in a formula unit.

↳ Not true - $(\text{NH}_4)_3\text{PO}_4$ - ammonium is a polyatomic cation.
 (NH_4^+)

23. Write chemical formulas for the following ternary ionic compounds.

- a. Calcium sulfate



- b. Copper(II) nitrate



- c. Lithium phosphate



- d. Potassium permanganate



- e. Aluminum sulfite



- f. Magnesium bicarbonate



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Rachel Turoscy



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