

Bond Energy

What makes a reaction endothermic or exothermic?

Why?

Chemical reactions can be a lot like playing with Legos®—you must take apart part of your last creation before you can replace it with something new. For many chemical reactions, we have to first break bonds in the reactants before we can put the atoms back into a new arrangement to form the products. Both of these processes involve changes in energy. The net energy change for a reaction is called the **heat of reaction** or the **change in enthalpy** (ΔH). In this activity we will look at one way energy changes can be approximated for chemical reactions.

Model 1 – Breaking and Forming Bonds

	Reaction	Change in Energy (kJ/mol _{rxn})		Reaction	Change in Energy (kJ/mol _{rxn})
A)	$\text{PCl}_3(\text{g}) \rightarrow \text{P}(\text{g}) + 3\text{Cl}(\text{g})$	+ 966.7 <i>broken</i>	E)	$\text{P}(\text{g}) + 3\text{Cl}(\text{g}) \rightarrow \text{PCl}_3(\text{g})$	- 966.7 <i>formed</i>
B)	$\text{PCl}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 5\text{Cl}(\text{g})$	+ 1,297.9 <i>broken</i>	F)	$\text{P}(\text{g}) + 5\text{Cl}(\text{g}) \rightarrow \text{PCl}_5(\text{g})$	- 1,297.9 <i>formed</i>
C)	$\text{PF}_3(\text{g}) \rightarrow \text{P}(\text{g}) + 3\text{F}(\text{g})$	+ 1,470.4 <i>broken</i>	G)	$\text{P}(\text{g}) + 3\text{F}(\text{g}) \rightarrow \text{PF}_3(\text{g})$	- 1,470.4 <i>formed</i>
D)	$\text{PF}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 5\text{F}(\text{g})$	+ 2,305.4 <i>broken</i>	H)	$\text{P}(\text{g}) + 5\text{F}(\text{g}) \rightarrow \text{PF}_5(\text{g})$	- 2,305.4 <i>formed</i>

- Identify four reactions in Model 1 where bonds are being broken. Write "broken" in the Change in Energy box in the table for each reaction you identify.
- Identify four reactions in Model 1 where bonds are being formed. Write "formed" in the Change in Energy box in the table for each reaction you identify.
- Circle the correct word to complete each sentence below.
 - When bonds are (broken/formed) there is a positive energy change.
 - Breaking bonds is (endothermic/exothermic).
 - When bonds are (broken/formed) there is a negative energy change.
 - Forming bonds is (endothermic/exothermic).
- Find two reactions in Model 1 that are exact opposites of each other, that is, one reaction is the reverse of the other reaction.

- How do the changes in energy for the reverse reactions compare?

The changes in energy for these reactions are the same, but opposite in sign.

- Explain your answer to part a considering what you learned from Questions 1–3 about bond breaking and bond formation.

If a given amount of energy is required to break a bond, then the same amount of energy will be released when the bond is formed.

Pogil Activities For High School Chemistry Bond Energy Answers

Terry C. Jones



Pogil Activities For High School Chemistry Bond Energy Answers:

Enjoying the Tune of Appearance: An Psychological Symphony within **Pogil Activities For High School Chemistry Bond Energy Answers**

In a world taken by monitors and the ceaseless chatter of instantaneous connection, the melodic elegance and psychological symphony created by the prepared word usually diminish into the backdrop, eclipsed by the persistent sound and distractions that permeate our lives. However, situated within the pages of **Pogil Activities For High School Chemistry Bond Energy Answers** a stunning fictional treasure filled with raw feelings, lies an immersive symphony waiting to be embraced. Crafted by a wonderful composer of language, that charming masterpiece conducts viewers on a mental journey, well unraveling the hidden melodies and profound impact resonating within each cautiously constructed phrase. Within the depths of the moving assessment, we will investigate the book is key harmonies, analyze their enthralling publishing type, and surrender ourselves to the profound resonance that echoes in the depths of readers souls.

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Table of Contents Pogil Activities For High School Chemistry Bond Energy Answers

1. Understanding the eBook Pogil Activities For High School Chemistry Bond Energy Answers
 - The Rise of Digital Reading Pogil Activities For High School Chemistry Bond Energy Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Pogil Activities For High School Chemistry Bond Energy Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pogil Activities For High School Chemistry Bond Energy Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pogil Activities For High School Chemistry Bond Energy Answers

- Personalized Recommendations
 - Pogil Activities For High School Chemistry Bond Energy Answers User Reviews and Ratings
 - Pogil Activities For High School Chemistry Bond Energy Answers and Bestseller Lists
5. Accessing Pogil Activities For High School Chemistry Bond Energy Answers Free and Paid eBooks
 - Pogil Activities For High School Chemistry Bond Energy Answers Public Domain eBooks
 - Pogil Activities For High School Chemistry Bond Energy Answers eBook Subscription Services
 - Pogil Activities For High School Chemistry Bond Energy Answers Budget-Friendly Options
 6. Navigating Pogil Activities For High School Chemistry Bond Energy Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Pogil Activities For High School Chemistry Bond Energy Answers Compatibility with Devices
 - Pogil Activities For High School Chemistry Bond Energy Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pogil Activities For High School Chemistry Bond Energy Answers
 - Highlighting and Note-Taking Pogil Activities For High School Chemistry Bond Energy Answers
 - Interactive Elements Pogil Activities For High School Chemistry Bond Energy Answers
 8. Staying Engaged with Pogil Activities For High School Chemistry Bond Energy Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pogil Activities For High School Chemistry Bond Energy Answers
 9. Balancing eBooks and Physical Books Pogil Activities For High School Chemistry Bond Energy Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pogil Activities For High School Chemistry Bond Energy Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Pogil Activities For High School Chemistry Bond Energy Answers
 - Setting Reading Goals Pogil Activities For High School Chemistry Bond Energy Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Pogil Activities For High School Chemistry Bond Energy Answers

- Fact-Checking eBook Content of Pogil Activities For High School Chemistry Bond Energy Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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