

The Nucleus & Radioactive Decay – Ch. 22 (p.701-712)

PART A – MASS DEFECT & NUCLEAR BINDING ENERGY

1. Find the mass defect of oxygen-16 if the measured atomic mass is 15.994915 amu.
1 proton = 1.007276 amu, 1 neutron = 1.008665 amu, and 1 electron = 0.0005486 amu.

2. Calculate the nuclear binding energy of oxygen-16 given that 1 amu = 1.6605×10^{-27} kg.

PART B – NUCLEAR DECAY

Write equations for the following nuclear decay reactions. Make sure that both mass numbers and atomic numbers are balanced on each side.

- | | |
|--|---|
| 3. Decay of polonium-218 by alpha (α) emission. | 5. Decay of chlorine-32 by positron (β^+) emission. |
| 4. Decay of carbon-14 by beta (β^-) emission. | 6. Decay of promethium-142 by electron capture. |

PART C – HALF-LIFE

7. The half-life of phosphorous-30 is 2.5 min. If you start with 35 g of phosphorus-30, how many grams would remain after 20.0 min?

8. The half-life of polonium-210 is 138.4 days. How many milligrams of polonium-210 remain after 415.2 days if you start with 2.0 mg of the isotope?

9. 20.0 g of a radioactive isotope are present at 1:00 p.m., and 5.0 g remain at 2:00 p.m.
 - a. How many half-lives have gone by? _____
 - b. How long is the half-life of the isotope? _____
 - c. Predict how many grams will be left at 2:30 p.m. _____

Radioactive Decay Worksheet D Answers

Léo Pomar



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