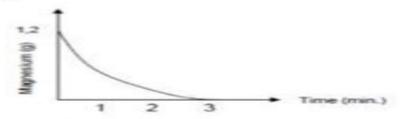
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records the change in the mass of magnesium as the reaction proceeds at regular intervals. The balanced chemical equation for the reaction is:

$$Mg(s) + 2HCl(aq) \rightarrow MgCl_2(aq) + H_2(g)$$

The change in the mass of magnesium during the reaction is shown on the graph below.



- 6.2.1 Identify the limiting agent in this reaction. Give a reason for the answer.
- 6.2.2 Calculate the number of moles of unreacted hydrochloric acid in the test tube after 3 minutes.

# **Hide Answers**

#### QUESTION/VRAAG 6

- 6.1.1 2H<sub>2</sub>O<sub>2</sub> (aq) ✓ → 2H<sub>2</sub>O (ℓ) + O<sub>2</sub> (g) ✓
- 6.1.2 The catalyst lowers the activation energy of the reaction 
  Accept: catalyst speeds up the reaction
  In Katalisator vertiage die aktiveringsenergie van die reaksie
  - 'n Katalisator verlaag die aktiveringsenergie van die reaksie 🗸 🗸 Aanvaar: katalisator laat die reaksie vinniger plaasvind
- 6.1.3 OPTION 1/OPSIE 1

$$0,049 = \frac{m}{34} \checkmark$$

OPTION 2/OPSIE 2
From the balanced equation:
Vanaf gebalanseerde
vergelyking:

68g  $H_2O_2 \rightarrow 24,45 \text{ dm}^3 O_2 \checkmark \checkmark$   $X \text{ g } H_2O_2 \rightarrow 600 \times 10^{-3} \text{ dm}^3 \checkmark$  $X = 68 \times 0.6 \checkmark$  (2)

(7) [19]

(2)

$$X = \frac{68 \times 0.6}{24.45}$$
   
  $X = 1.67 g$ 

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- 1. Understanding the eBook Physical Science Grade 11 Exampler September 2014
  - The Rise of Digital Reading Physical Science Grade 11 Exampler September 2014
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Physical Science Grade 11 Exampler September 2014
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - o Features to Look for in an Physical Science Grade 11 Exampler September 2014
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Physical Science Grade 11 Exampler September 2014
  - Personalized Recommendations
  - Physical Science Grade 11 Exampler September 2014 User Reviews and Ratings
  - Physical Science Grade 11 Exampler September 2014 and Bestseller Lists
- 5. Accessing Physical Science Grade 11 Exampler September 2014 Free and Paid eBooks
  - Physical Science Grade 11 Exampler September 2014 Public Domain eBooks
  - Physical Science Grade 11 Exampler September 2014 eBook Subscription Services
  - Physical Science Grade 11 Exampler September 2014 Budget-Friendly Options

- 6. Navigating Physical Science Grade 11 Exampler September 2014 eBook Formats
  - o ePub, PDF, MOBI, and More
  - Physical Science Grade 11 Exampler September 2014 Compatibility with Devices
  - Physical Science Grade 11 Exampler September 2014 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - o Adjustable Fonts and Text Sizes of Physical Science Grade 11 Exampler September 2014
  - Highlighting and Note-Taking Physical Science Grade 11 Exampler September 2014
  - Interactive Elements Physical Science Grade 11 Exampler September 2014
- 8. Staying Engaged with Physical Science Grade 11 Exampler September 2014
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Physical Science Grade 11 Exampler September 2014
- 9. Balancing eBooks and Physical Books Physical Science Grade 11 Exampler September 2014
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Physical Science Grade 11 Exampler September 2014
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science Grade 11 Exampler September 2014
  - Setting Reading Goals Physical Science Grade 11 Exampler September 2014
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science Grade 11 Exampler September 2014
  - Fact-Checking eBook Content of Physical Science Grade 11 Exampler September 2014
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
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