

Balancing Equations Answers

1. $2 \text{H}_2 + \text{O}_2 \rightarrow 2 \text{H}_2\text{O}$
2. $\text{H}_3\text{PO}_4 + 3 \text{KOH} \rightarrow \text{K}_3\text{PO}_4 + 3 \text{H}_2\text{O}$
3. $6 \text{K} + \text{B}_2\text{O}_3 \rightarrow 3 \text{K}_2\text{O} + 2 \text{B}$
4. $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$
5. $10 \text{Na} + 2 \text{NaNO}_3 \rightarrow 6 \text{Na}_2\text{O} + \text{N}_2$
6. $4 \text{C} + \text{S}_8 \rightarrow 4 \text{CS}_2$
7. $2 \text{Na} + \text{O}_2 \rightarrow \text{Na}_2\text{O}_2$
8. $2 \text{N}_2 + 5 \text{O}_2 \rightarrow 2 \text{N}_2\text{O}_5$
9. $2 \text{H}_3\text{PO}_4 + 3 \text{Mg(OH)}_2 \rightarrow \text{Mg}_3(\text{PO}_4)_2 + 6 \text{H}_2\text{O}$
10. $2 \text{NaOH} + \text{H}_2\text{CO}_3 \rightarrow \text{Na}_2\text{CO}_3 + 2 \text{H}_2\text{O}$
11. $\text{KOH} + \text{HBr} \rightarrow \text{KBr} + \text{H}_2\text{O}$
12. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}_2$
13. $4 \text{Na} + \text{O}_2 \rightarrow 2 \text{Na}_2\text{O}$
14. $2 \text{Al(OH)}_3 + 3 \text{H}_2\text{CO}_3 \rightarrow \text{Al}_2(\text{CO}_3)_3 + 6 \text{H}_2\text{O}$
15. $16 \text{Al} + 3 \text{S}_8 \rightarrow 8 \text{Al}_2\text{S}_3$
16. $6 \text{Cs} + \text{N}_2 \rightarrow 2 \text{Cs}_3\text{N}$
17. $\text{Mg} + \text{Cl}_2 \rightarrow \text{MgCl}_2$
18. $10 \text{Rb} + 2 \text{RbNO}_3 \rightarrow 6 \text{Rb}_2\text{O} + \text{N}_2$
19. $2 \text{C}_6\text{H}_6 + 15 \text{O}_2 \rightarrow 12 \text{CO}_2 + 6 \text{H}_2\text{O}$
20. $\text{N}_2 + 3 \text{H}_2 \rightarrow 2 \text{NH}_3$
21. $2 \text{C}_{10}\text{H}_{22} + 31 \text{O}_2 \rightarrow 20 \text{CO}_2 + 22 \text{H}_2\text{O}$
22. $\text{Al(OH)}_3 + 3 \text{HBr} \rightarrow \text{AlBr}_3 + 3 \text{H}_2\text{O}$
23. $2 \text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3 + 13 \text{O}_2 \rightarrow 8 \text{CO}_2 + 10 \text{H}_2\text{O}$
24. $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$
25. $\text{C}_3\text{H}_8 + 5 \text{O}_2 \rightarrow 3 \text{CO}_2 + 4 \text{H}_2\text{O}$
26. $3 \text{Li} + \text{AlCl}_3 \rightarrow 3 \text{LiCl} + \text{Al}$
27. $2 \text{C}_2\text{H}_6 + 7 \text{O}_2 \rightarrow 4 \text{CO}_2 + 6 \text{H}_2\text{O}$
28. $3 \text{NH}_4\text{OH} + \text{H}_3\text{PO}_4 \rightarrow (\text{NH}_4)_3\text{PO}_4 + 3 \text{H}_2\text{O}$
29. $3 \text{Rb} + \text{P} \rightarrow \text{Rb}_3\text{P}$
30. $\text{CH}_4 + 2 \text{O}_2 \rightarrow \text{CO}_2 + 2 \text{H}_2\text{O}$
31. $2 \text{Al(OH)}_3 + 3 \text{H}_2\text{SO}_4 \rightarrow \text{Al}_2(\text{SO}_4)_3 + 6 \text{H}_2\text{O}$
32. $2 \text{Na} + \text{Cl}_2 \rightarrow 2 \text{NaCl}$
33. $16 \text{Rb} + \text{S}_8 \rightarrow 8 \text{Rb}_2\text{S}$
34. $2 \text{H}_3\text{PO}_4 + 3 \text{Ca(OH)}_2 \rightarrow \text{Ca}_3(\text{PO}_4)_2 + 6 \text{H}_2\text{O}$
35. $\text{NH}_3 + \text{HCl} \rightarrow \text{NH}_4\text{Cl}$
36. $2 \text{Li} + 2 \text{H}_2\text{O} \rightarrow 2 \text{LiOH} + \text{H}_2$
37. $\text{Ca}_3(\text{PO}_4)_2 + 3 \text{SiO}_2 + 5 \text{C} \rightarrow 3 \text{CaSiO}_3 + 5 \text{CO} + 2 \text{P}$
38. $4 \text{NH}_3 + 3 \text{O}_2 \rightarrow 2 \text{N}_2 + 6 \text{H}_2\text{O}$
39. $4 \text{FeS}_2 + 11 \text{O}_2 \rightarrow 2 \text{Fe}_2\text{O}_3 + 8 \text{SO}_2$
40. $5 \text{C} + 2 \text{SO}_2 \rightarrow \text{CS}_2 + 4 \text{CO}$
41. $\text{Fe} + \text{S} \rightarrow \text{FeS}$
42. $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$
43. $4 \text{Al} + 3 \text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$
44. $\text{Fe}_2\text{O}_3 + 3 \text{C} \rightarrow 3 \text{CO} + 2 \text{Fe}$
45. $\text{K}_2\text{O} + \text{H}_2\text{O} \rightarrow 2 \text{KOH}$
46. $\text{K}_2\text{CO}_3 + \text{BaCl}_2 \rightarrow 2 \text{KCl} + \text{BaCO}_3$
47. $\text{Mg(OH)}_2 + \text{H}_2\text{SO}_4 \rightarrow \text{MgSO}_4 + 2 \text{H}_2\text{O}$
48. $2 \text{KF} + \text{BaBr}_2 \rightarrow \text{BaF}_2 + 2 \text{KBr}$
49. $\text{HCl} + \text{NH}_3 \rightarrow \text{NH}_4\text{Cl}$
50. $\text{Bi}_2(\text{SO}_4)_3 + 6 \text{NH}_4\text{OH} \rightarrow 2 \text{Bi(OH)}_3 + 3 (\text{NH}_4)_2\text{SO}_4$

Physical Science Balancing 1 Answer Key

Wolfgang Guggemos

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Science, eBook Beth Barber,2004-09-01 This book supplements and enriches classroom teaching to enhance students understanding of vocabulary functions and fundamental processes of physical sciences work Topics include force and motion chemistry atoms and elements scientific process simple machines energy light and sound magnetism and electricity

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Mathematical Modeling in Physical Sciences Dimitrios Vlachos,2024-05-23 This volume gathers selected papers presented at the ICMSQUARE 2023 12th International Conference on Mathematical Modeling in Physical Sciences held in Belgrade Serbia from August 28 31 2023 This proceedings offers a compilation of cutting edge research which aims to advance the knowledge and development of high quality research in mathematical fields related to physics chemistry biology medicine economics environmental sciences and more Annually held since 2012 the ICMSQUARE conference serves as a platform for the exchange of ideas and

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