

## Personal Project Bibliography

Prepared by: [Your Name]  
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Department: [Your Department]

### I. Books

- **Smith, John.** *The Art of Research: A Practical Guide.* Publisher, 2057.  
*Annotation:* An insightful exploration of research methods and strategies for conducting personal projects. Smith offers practical advice and examples to guide researchers through the process.
- **Johnson, Emily.** *Creative Sparks: Nurturing Your Imagination.* Publisher, 2059.  
*Annotation:* Johnson's book provides inspiration and techniques for fostering creativity in personal projects. Her insights into overcoming creative blocks and finding innovative ideas are invaluable.

### II. Journal Articles

- **Brown, Sarah.** "The Power of Personal Projects." *Journal of Creative Endeavors*, vol. 15, no. 2050, Year, pp. 45-62.  
*Annotation:* Brown's article examines the significance of personal projects in fostering personal growth and development. She discusses how engaging in such endeavors can enhance skills and promote self-expression.
- **Lee, David.** "Navigating Online Research: Tips and Tools." *Digital Explorer*, vol. 8, no. 3, 2053, pp. 28-36.  
*Annotation:* Lee offers practical advice and tools for conducting effective online research. His article covers topics such as evaluating sources, using search engines efficiently, and avoiding plagiarism.

# Personal Project Paper Examples

**RJ Alexander**



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web write down the equilibrium equation you are investigating using the information provided by the introduction within the virtual lab activity equilibrium reaction  $CoH_2O + 6Cl_2 \rightleftharpoons CoCl_4 + 6H_2O$  select each control test tube to compare the concentrations of

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web the study of reaction rates will lead you into the study of chemical equilibrium in this semester you will also study electrochemistry from batteries to electroplating and you will learn about acids bases and salts unit 15 is a brief study of the atomic nucleus which your teacher may introduce at any time during the year

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web equilibrium moles  $HSCN$   $2.00 \times 10^{-5}$  mol  $3.00 \times 10^{-6}$  mol  $1.70 \times 10^{-5}$  mol  $HSCN$  knowing the number of moles of  $Fe^{3+}$  and  $HSCN$  present in the equilibrium mixture and the volume of the mixture we can easily find the concentrations of those two species mol  $Fe^{3+}$   $3.170 \times 10^{-5}$  mol  $Fe^{3+}$   $8.50 \times 10^{-4}$  M

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