

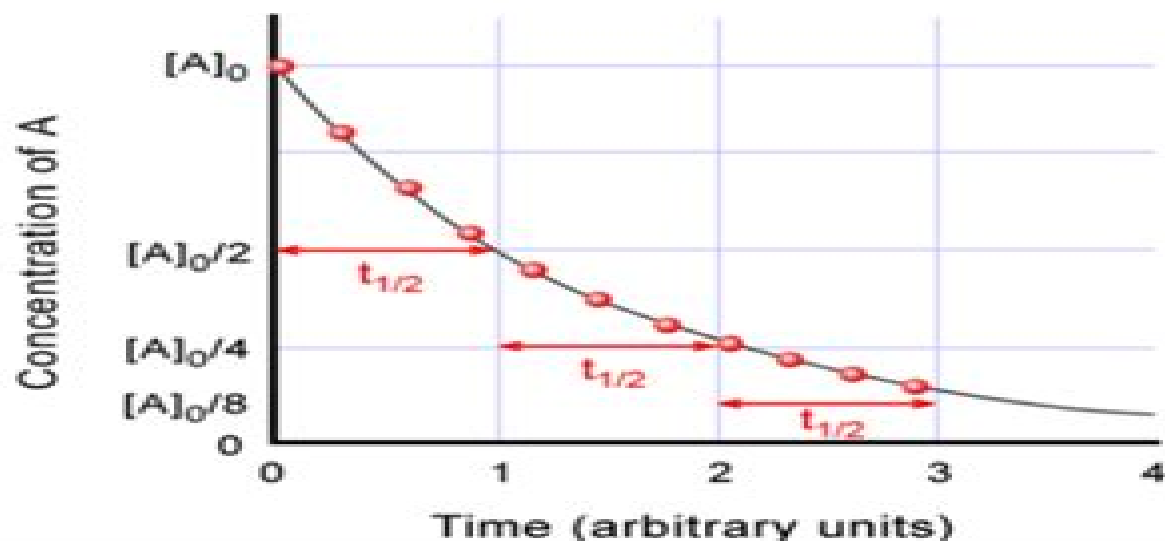
Penny Decay

Radioactive decay follows 1st order kinetics and in the reaction, the concentration of the reactant decreases exponentially. The rate of the reaction equals the concentration of the reactant, $[A]$, raised to the first power times a proportionality constant, k , which is called the rate constant. The rate constant is a fixed value for a given reaction.

$$\text{Rate} = k [A]$$

$[A]$ is the concentration of the reactant and k is the rate constant.

A plot of the concentration of A versus time is shown below. You can see that the concentration of A does not decrease in a linear manner. The concentration rapidly decreases at the beginning but slows down over time. This is characteristic of an exponential decrease.



Radioactive Decay Penny Lab Graph

**Alison Warr, Catherine Curtis, Penny
Slingerland**

A red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the authors' names.

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Chemistry in the Community. American Chemical Society, 2002 This volume has relevance to a wide number of courses giving a hands on introduction to chemistry in relation to community issues rather than around specific chemical concepts
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