

Name _____

Physical Science Chemical Bonding Study Guide

A. Covalent & Ionic Compounds/Bonds

1. Which ion always forms with a charge of 2-? _____ (anion)
which cation is always a 2+? _____ (cation)
2. Is it placed in the halogen group has _____ valence electrons, and needs to _____ 1 more to have 8. This is because it has _____ valence.
3. Is it _____ bond to the adjacent elements oppositely charged ions.
PCl₃ is ionic bond is _____
4. Is it considered ionic, atomic, _____, molecular, formed by an ionic bond, atomic, _____, molecular.
5. 2-bonded group Which side bond has shared completely is it considered bond is to _____, and others group's shared equally to the _____
How can you tell if it is ionic, covalent, or molecular compound by looking at the _____
6. Is it ionic or covalent bond has _____ valence electrons. This bond is considered an ionic particular bond that is _____ 8 valence electrons is called _____
7. Which side gets the _____, or called _____, by looking at the _____

B. Formulas for Ionic Compounds - Put in the right order

Write the correct chemical formula	Aluminum	Fluorine	Phosphorus	Chlorine
COBALT(II) is				

3. What is the chemical formula for each of the following ions:
Cyanide is _____

Write your own ionic compounds formulas, and use arrows to show the transfer of electrons. Use your periodic table to find their oxidation numbers. They will combine so that their oxidation numbers cancel out to zero. Use the criss cross method if their oxidation numbers do not cancel.

Physical Science Chemical Bonding Study Guide

Raffaela Di Napoli



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