



NATIONAL
SENIOR CERTIFICATE

GRADE 11

NOVEMBER 2020

PHYSICAL SCIENCES P2
(CHEMISTRY)
EXEMPLAR

DATE: / /
TIME: hours



The question paper consists of 10 pages, including 4 extra sheets.
Use blue or black ink for writing. Do not use correction fluid.

QUESTION 3

PHYSICAL SCIENCES PAPER 2

INTERMOLECULAR FORCES

GRADE 11

NOVEMBER 2020

Examiner's use: _____ Physical Sciences P2

Duration: 1 hour 45 minutes

3.1 The boiling points in the table below were obtained during an investigation carried out to measure the boiling points of pure 1-hexanol. The table also lists the ΔH_{vap} and ΔH_{fusion} for 1-hexanol in kJ mol⁻¹.

Compound	Molecular mass (M _r)	Boiling point (°C)
A	102	16,2
B	102	21

- 3.1.1 Define the term boiling point. (2)
- 3.1.2 Is pure 1-hexanol only composed of A and B? or A, B and C? or all of them present? (2)
- 3.1.3 Name the type of intermolecular force that exists between molecules of both compounds A and B and its boiling point of these molecules. (2)
- 3.1.4 What effect, if any, does compound C have on the boiling point of pure 1-hexanol? (2)
- 3.1.5 What compound (A or B) will have the higher vapour pressure? (2)
- 3.1.6 What compound (A or B) will have the higher surface tension? (2)

Compound C (M_r 102) will show in the table, has a smaller molecular mass than both compounds A and B but has a relatively higher boiling point of 16,2 °C.

- 3.1.6 Explain why the boiling point of compound C is 16,2 °C (higher than that of compound A and B) by referring to the ΔH_{vap} and intermolecular or intramolecular forces involved. (2)

3.2 Add column C to the table in Table 1, showing the boiling point of pure 1-hexanol (C₆H₁₄) and water. (2)

1-hexanol	Water
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- 3.2.1 Is water more volatile than 1-hexanol? (2)
- 3.2.2 Explain the process in QUESTION 3.1.1 about the boiling in the intermolecular forces involved. (2)

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SA Adler



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