

Gas  
Liquid  
Solid  
Unknown

# Periodic Table

Cheat Sheet

Reactions

Single Displacement  $\rightarrow A + BC = B + AC$   
Double Displacement  $\rightarrow AB + CD = BC + AD$   
Synthesis  $\rightarrow A + B = AB$   
Decomposition  $\rightarrow AB = A + B$   
Combustion  $\rightarrow \text{Hydrocarbon} + O_2 = CO_2 + H_2O$   
Neutralization  $\rightarrow \text{Acid} + \text{Base} = \text{Salt} + H_2O$

Polyatomic ions (ion composed of multiple atoms)

$NO_3^- \rightarrow$  Nitrate (-1)

$PO_4^{3-} \rightarrow$  Phosphate (-3)

$OH^- \rightarrow$  Hydroxide (-1)

$HCO_3^- \rightarrow$  Bicarbonate (-1)

$NO_2^- \rightarrow$  Nitrite (-1)

$SO_4^{2-} \rightarrow$  Sulfate (-2)

$CO_3^{2-} \rightarrow$  Carbonate (-2)

$NH_4^+ \rightarrow$  Ammonium (+1)

H	He																		
Li	Be																		
Na	Mg																		
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr		
Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe		
Cs	Ba	57-71	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn		
Fr	Ra	89-103	Rf	Db	Sg	Bh	Hs	Mt	Ds	Rg	Cn	Nh	Fl	Mc	Lv	Ts	Og		

**Alkali Metals:**  
One valence electron  
Very reactive because they are ready to lose an electron to form cations  
Weakly in ionic bonds due to reactivity  
Bonds with halogens, and other non-metals  
Have a vigorous reaction when bonding with halogens, oxygen and the compound of water

**Alkaline Earth Metals:**  
Two valence electrons  
Inertive but not as reactive as halogens and alkali metals because it needs to lose 2 electrons to have a full valence shell  
Bonds with halogens, and other non-metals

**Transition Metals/Transition Metals:** Transition metals are metals that vary in valence electrons and can have multiple ion charges. They have similar qualities as alkali metals and alkaline earth metals; they can form metallic and ionic bonds. Metals with multiple ion charges are called multivalent metals. For example, copper has an ion charge of +1 and +2 so it is a multivalent metal. Therefore, the ion charge must be specified when written in word form for ionic bonds with multivalent metals. Metals can also bond with polyatomic ions.

La	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu
Ac	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr

**Metalloids:** Contain metal and non-metal properties. They react with metals and non-metals to form covalent and ionic bonds but are less reactive than non-metals, depending on the electronegativity of the atoms.

**Non-metals:** They can not conduct electricity because they can not give off electrons, so current will not be able to go through.  
Non-metals bond with metals in an ionic bond and other non-metals in a molecular bond. They can also bond with the polyatomic ion ammonium due to its +1 charge.

**Noble Gases:** Non-metals in a previous form that have a full valence shell. This makes it not reactive, so it can not bond with other non-metals and metals. This is why noble gases are never in compounds and chemical reactions.

# Periodic Table Of Elements Grade 10

**Lauren Gardner**



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