



Name: ANSWER KEY
Date: _____

THE PERIODIC TABLE

Review Worksheet

1. Fill in the blanks with the terms and words most appropriate to complete each sentence.

- a) In the mid-nineteenth century, only sixty or so elements had been discovered. In 1869 (a year) a Russian chemist and inventor named Dmitri Mendeleev created a chart called the periodic table that predicted the properties of elements that had not yet been discovered. He did this by arranging the elements according to their increasing atomic numbers and by their physical and chemical properties.
- b) Each element in the periodic table is represented by an element box. In this box there is the following information: two numbers, one is a whole number which is called the atomic number, and the other number is larger and contains a decimal, and this one is called the atomic mass. In the box, there is also the short form of the element called the element symbol, and then the full identity of the element which is written using the element name.
- c) Today the modern periodic table is arranged into columns and rows. The columns are called groups and there are 18 (a number) of them. The rows are called periods and there are 7 (a number) of them. The periodic table is divided by a zigzag line also known as the staircase that separates the metals, found on the left of it, from the non-metals, found on the right of it. Directly around the zigzag line are eight elements called metalloids.
- d) All the metals, except for mercury, which is a liquid, are found as solids at room temperature. The metals are also mostly shiny in color, are ductile, malleable, conduct electricity and some are attracted to magnets.
- e) At room temperature, most of the non-metals are found in the gas state except for bromine which is a liquid, and iodine selenium carbon, phosphorus and sulfur which are all found as solids. These elements are colorless or come in various colors, are not ductile, not malleable and do not conduct electricity nor are they attracted to magnets.
- f) Metalloids are elements that have some properties of metals and some properties of non-metals. They are all found in the solid state. Most have a silvery metallic luster, they are brittle and are medium to good electrical conductors.
- g) Chemical families are groups on the periodic table that contain elements that share similar properties.
- h) Group 1 elements are called the alkali metals and they contain the most reactive group of all the metals. They react very violently with water and therefore must be stored under mineral oil so that the water vapor in the air cannot react with these elements. These metals get more reactive as you go down (up or down) the group and the most reactive of all the elements in this group is francium.
- i) Group 2 elements are called the alkaline earth metals and they contain the second most reactive group of all the metals. They also react with water and get more reactive as you go down (up or down) the group.

Periodic Table Mixed Review Answer Key

Daniela Niemeyer



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Modern Chemistry Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001

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Brown, H. Eugene LeMay Jr., Bruce E. Bursten, Catherine Murphy, Patrick Woodward, Steven Langford, Dalius Sagatys, Adrian George, 2013-10-04 If you think you know the Brown LeMay Bursten Chemistry text think again In response to market request we have created the third Australian edition of the US bestseller Chemistry The Central Science An extensive revision has taken this text to new heights Triple checked for scientific accuracy and consistency this edition is a more seamless and cohesive product yet retains the clarity innovative pedagogy functional problem solving and visuals of the previous version All artwork and images are now consistent in quality across the entire text And with a more traditional and logical organisation of the Organic Chemistry content this comprehensive text is the source of all the information and practice problems students are likely to need for conceptual understanding development of problem solving skills reference and test preparation **Science Education for Teacher Trainees and In-service Teachers** R. M. KALRA, 2008-01-01

This book in the field of science education offers a modern approach to education and construction of the school science

curriculum It lays emphasis on the role of science in transforming the thinking and behaviour pattern of students The book explains the philosophy of the processes of science teaching with a focus on values as an integral part of the programme examination and evaluation in science education and generalizations regarding the learning processes and their implications for science education Topics such as methods of science teaching laboratory facilities objective based science curriculum development and interdisciplinary and integrated approach to science teaching at the school level are discussed in detail Besides the topics such as Action Research and Forgotten Silent Majority have also been incorporated to encourage excellence in science education among academics Key Features Focuses on innovative methods for science teaching Discusses science education in the context of globalization Includes interesting thought provoking questions at the end of each chapter to encourage group discussions This book is intended for the students undergoing elementary teacher training courses nursery teacher training courses and courses in B Ed B A Education and M A Education It will also be immensely helpful to in service science teachers for the effective teaching of science **Chemistry and Society** Michael E Green,2019

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