

CH 5.3 - Periodic Trends WebQuest PART I

There is a .pdf version of this assignment on my blog. Please utilize the online version to link the webpages in order to avoid typing in the addresses. Reminder: columns on the periodic table are called families or groups and rows are called periods

Atomic Radius (pages 174-175)

Using your text, please answer the following:

1. Can the radius of an atom be measured directly? _____
2. In what unit is atomic radius usually measured? _____
3. Define Atomic Radius: _____

Go to the following website:

http://www.webelements.com/periodicity/atomic_radius_empirical/bar_chart.html

4. Look at the graph. What is on the X Axis? _____
5. What is on the Y Axis (include units)? _____
6. **The colors on the graph represent rows on the periodic table.**
In a complete sentence, answer - What appears to be the trend in atomic radius as you move from left to right in a row?

In a complete sentence, answer - What appears to be the trend in atomic radius as you move down a column?

7. Using your text, please answer the following:

- a. What is the general trend for atomic size within a group/family from top to bottom?
- b. What two variables affect atomic size within a group?
- c. What is the general trend in atomic size from left to right within a period?

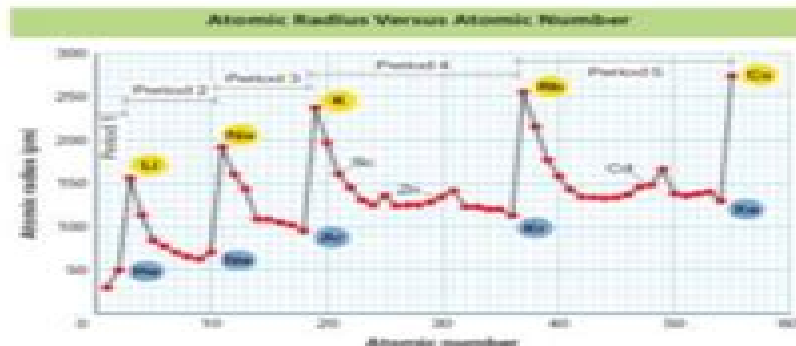
8. Use the graph to the right to answer the following:

- a. Determine which metal has an atomic radius of 238 pm?
- b. Is an atom of barium, atomic number 56, smaller or larger than an atom of cesium (Cs)?
- c. Which has a larger atom, helium or argon?
- d. Which has a larger atom, potassium or argon?

Teacher Checkpoint _____

Ions and Ionization Energy (pages 177-181)

Using your text, please answer the following:



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Lingjun Ying



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Table of Contents Periodic Trends Webquest Answers

1. Understanding the eBook Periodic Trends Webquest Answers
 - The Rise of Digital Reading Periodic Trends Webquest Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Webquest Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends Webquest Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Webquest Answers
 - Personalized Recommendations
 - Periodic Trends Webquest Answers User Reviews and Ratings

- Periodic Trends Webquest Answers and Bestseller Lists
- 5. Accessing Periodic Trends Webquest Answers Free and Paid eBooks
 - Periodic Trends Webquest Answers Public Domain eBooks
 - Periodic Trends Webquest Answers eBook Subscription Services
 - Periodic Trends Webquest Answers Budget-Friendly Options
- 6. Navigating Periodic Trends Webquest Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Trends Webquest Answers Compatibility with Devices
 - Periodic Trends Webquest Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trends Webquest Answers
 - Highlighting and Note-Taking Periodic Trends Webquest Answers
 - Interactive Elements Periodic Trends Webquest Answers
- 8. Staying Engaged with Periodic Trends Webquest Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends Webquest Answers
- 9. Balancing eBooks and Physical Books Periodic Trends Webquest Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends Webquest Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Trends Webquest Answers
 - Setting Reading Goals Periodic Trends Webquest Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Trends Webquest Answers
 - Fact-Checking eBook Content of Periodic Trends Webquest Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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