

# Percent Yield Practice

3 pages + 3-page Answer KEY + Periodic Table

Chemistry

## PERCENT YIELD

Show all work and circle or highlight your final answer.

$$\text{percent yield (\%)} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$$

- 1) 10.0 grams of sodium reacts with excess chlorine gas.



The theoretical yield of sodium chloride is 25.0 grams. The actual yield was 23.0 grams. Calculate the percent yield of sodium chloride.

Name \_\_\_\_\_

Date \_\_\_\_\_

- 2) Hydrogen gas ( $\text{H}_2$ ) reacts with oxygen gas ( $\text{O}_2$ ) to produce water ( $\text{H}_2\text{O}$ ). 5.0 grams of hydrogen reacts with excess oxygen gas.



What is the theoretical yield?

- 3) 4.0 grams of carbon reacts with excess oxygen gas.



The theoretical yield of carbon dioxide is 25.0 g. Calculate the percent yield of carbon dioxide.

- 4) 1.0 grams of zinc reacts with excess sulfur.



The theoretical yield of zinc sulfide is 15.0 grams. Calculate the percent yield of zinc sulfide.

- 5) The actual yield was 10.0 grams. Calculate the percent yield.

- 6) Magnesium ( $\text{Mg}$ ) reacts with hydrochloric acid ( $\text{HCl}$ ) to produce magnesium chloride ( $\text{MgCl}_2$ ) and hydrogen gas ( $\text{H}_2$ ). Calculate the theoretical yield.



What is the theoretical yield?

- 7) 1.0 grams of phosphorus reacts with excess chlorine gas.



The theoretical yield of phosphorus trichloride is 10.0 grams. Calculate the percent yield of phosphorus trichloride.

Name \_\_\_\_\_

Date \_\_\_\_\_

- 8) Calcium carbonate ( $\text{CaCO}_3$ ) reacts with hydrochloric acid ( $\text{HCl}$ ) to produce calcium chloride ( $\text{CaCl}_2$ ), carbon dioxide ( $\text{CO}_2$ ), and water ( $\text{H}_2\text{O}$ ). 10.0 grams of calcium carbonate reacts with excess hydrochloric acid.



What is the theoretical yield of calcium chloride?

$$\begin{array}{r} 10.0 \text{ g CaCO}_3 \times \frac{1 \text{ mol CaCO}_3}{100.09 \text{ g CaCO}_3} \times \frac{1 \text{ mol CaCl}_2}{1 \text{ mol CaCO}_3} \times \frac{110.98 \text{ g CaCl}_2}{1 \text{ mol CaCl}_2} = 11.0 \text{ g CaCl}_2 \end{array}$$

$$\begin{array}{l} \text{Molar Mass CaCl}_2 \\ \text{Ca} = 40.08 \\ \text{Cl} = 35.45 \\ 2 \times 35.45 = 70.90 \\ 40.08 + 70.90 = 110.98 \text{ g} \end{array}$$

$$\begin{array}{r} 10.0 \text{ g CaCO}_3 \times \frac{1 \text{ mol CaCl}_2}{100.09 \text{ g CaCO}_3} \times \frac{1 \text{ mol CaCl}_2}{1 \text{ mol CaCO}_3} \times \frac{110.98 \text{ g CaCl}_2}{1 \text{ mol CaCl}_2} = 11.0 \text{ g CaCl}_2 \end{array}$$

$$\begin{array}{l} \text{Molar Mass CaCl}_2 \\ \text{Ca} = 40.08 \\ \text{Cl} = 35.45 \times 2 = 70.90 \\ 40.08 + 70.90 = 110.98 \text{ g} \end{array}$$

- 9) The actual yield was 11.0 grams. Calculate the percent yield of calcium chloride.

$$\text{percent yield (\%)} = \frac{11.0 \text{ g}}{11.0 \text{ g}} \times 100 = 100\%$$

- 10) Potassium ( $\text{K}$ ) reacts with iodine ( $\text{I}_2$ ) to produce potassium iodide ( $\text{KI}$ ). 10.0 grams of potassium reacts with excess iodine.



What is the theoretical yield?

$$\begin{array}{r} 10.0 \text{ g K} \times \frac{1 \text{ mol K}}{39.10 \text{ g K}} \times \frac{2 \text{ mol KI}}{2 \text{ mol K}} \times \frac{166.10 \text{ g KI}}{1 \text{ mol KI}} = 16.6 \text{ g KI} \end{array}$$

$$\begin{array}{l} \text{Molar Mass KI} \\ \text{K} = 39.10 \\ \text{I} = 126.90 \\ 39.10 + 126.90 = 166.10 \text{ g} \end{array}$$

ANSWER KEY INCLUDED

YIELD

# Percent Yield Example Answer Key

**Christian Drosten**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

## **Percent Yield Example Answer Key:**

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