## Finding the Percent Composition

Percent Composition - The percentage (by mass) of each element in a compound.

#### Example 1:

Find the percent composition of red candies in a bag of 200 candies, if there are 26 red, and 174 blue candies.

Total Candies: 200

Red Candies: 26

Blue Candies: 174

26 red	× 100	
200 total	A IUU	

Formula:

Part x 100 Total

#### Example 2:

Find the percent composition of MgCl<sub>2</sub>.

- Total mass MgCl2: 95.21 g
- Mass of Mg: 24.31 g
- 70.91 g Mass of Cl2:

#### Example 3:

Find the percent composition of H2SO4.

- Total mass H<sub>2</sub>SO<sub>4</sub>: 98.08 g
- 2.02 g Mass of H<sub>2</sub>:
- Mass of S: 32.07 g
- Mass of O4: 64.00 g

$$\frac{2.02 \text{ g}}{98.08 \text{ g}} \times 100$$

$$\frac{2.02 \,\mathrm{g}}{98.08 \,\mathrm{g}} \times 100 = 2.06 \,\% \,\mathrm{H}_2$$

# **Percent Composition Answers**

**Victor Gage Kimball** 

**Percent Composition Answers:** 

#### **Percent Composition Answers** Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has be much more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Percent Composition Answers**," written by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we shall delve into the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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