



Nokia Lumia 635 User Manual

Michael Rajnik



Nokia Lumia 635 User Manual:

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Blackberry It s also for those who need to broaden their skillset by adding more data extraction and recovery techniques What You Will Learn Retrieve mobile data using modern forensic tools Work with Oxygen Forensics for Android devices acquisition Perform a deep dive analysis of iOS Android Windows and BlackBerry Phone file systems Understand the importance of cloud in mobile forensics and extract data from the cloud using different tools Learn the application of SQLite and Plists Forensics and parse data with digital forensics tools Perform forensic investigation on iOS Android Windows and BlackBerry mobile devices Extract data both from working and damaged mobile devices using JTAG and Chip off Techniques In Detail Considering the emerging use of mobile phones there is a growing need for mobile forensics Mobile forensics focuses specifically on performing forensic examinations of mobile devices which involves extracting recovering and analyzing data for the purposes of information security criminal and civil investigations and internal investigations Mobile Forensics Cookbook starts by explaining SIM cards acquisition and analysis using modern forensics tools You will discover the different software solutions that enable digital forensic examiners to quickly and easily acquire forensic images You will also learn about forensics analysis and acquisition on Android iOS Windows Mobile and BlackBerry devices Next you will understand the importance of cloud computing in the world of mobile forensics and understand different techniques available to extract data from the cloud Going through the fundamentals of SQLite and Plists Forensics you will learn how to extract forensic artifacts from these sources with appropriate tools By the end of this book you will be well versed with the advanced mobile forensics techniques that will help you perform the complete forensic acquisition and analysis of user data stored in different devices Style and approach This book delivers a series of extra techniques and methods for extracting and analyzing data from your Android iOS Windows and Blackberry devices Using practical recipes you will be introduced to a lot of modern forensics tools for performing effective mobile forensics

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web limiting reactant theoretical and percentage yields key 4 $\text{ko}_2 + 2 \text{h}_2\text{o} \rightarrow 4 \text{koh} + \text{o}_2$ g if a reaction vessel contains 0 mol ko_2 and 0 mol h_2o what is the limiting reactant

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$$4\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Cu} + 2\text{Ag} + 2\text{CuNO}_3 + 2\text{AgNO}_3$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole H₂S and multiply it with 2 mole H₂O over 2 mole H₂S from the equation

$$8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8.3 \text{ mole H}_2\text{O}$$

8.3 mole H₂S x 18g H₂O = 149.4g H₂O is your theoretical yield so percent yield = $\frac{137.1\text{g}}{149.4\text{g}} \times 100 = 91.77\%$

problems 6 & 8 are similar to problem 5

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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$

if the reaction of 21.3 grams of O₂ produces a 47.3% yield how many grams of CO would be produced

step one figure out how many moles of the limiting reagent you have

$$21.3 \text{ g O}_2 \times \frac{1 \text{ mole}}{32 \text{ g}} = 0.666 \text{ moles O}_2$$

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web percent yield is a measure of the efficiency of the experimental design

$$\text{yield} = \frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$$

in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction

$$4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$$

b determine the theoretical yield of Al₂O₃

$$160.0 \text{ g Al} \times \frac{1 \text{ mole Al}}{27 \text{ g}} \times \frac{2 \text{ mole Al}_2\text{O}_3}{4 \text{ mole Al}} \times 101.96 \text{ g Al}_2\text{O}_3 = 150.1 \text{ g Al}_2\text{O}_3$$

c determine the percent yield

$$\frac{126.0 \text{ g}}{150.1 \text{ g}} \times 100 = 83.9\%$$

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