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Emma Taylor,JK. Hudson



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web a reaction and its experimental rate law are represented below $2 \text{ICl} + \text{H}_2 \text{g} \rightarrow 2 \text{HCl} + \text{I}_2 \text{g}$ rate $k [\text{ICl}]^2 [\text{H}_2]$ a chemist proposes a mechanism for the reaction that is consistent with the rate law the mechanism has two elementary steps and the first step is slow compared to the second

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web mar 13 2023 work in groups on these problems you should try to answer the questions without referring to your textbook if you get stuck try asking another group for help 1 for the reaction below if substance a is disappearing at a rate of $1.82 \times 10^{-2} \text{ mol l}^{-1} \text{ s}^{-1}$ at what rate is c appearing 3 a 3 b $\rightarrow$ 5 c 2 d

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consider the following mechanism
 $\text{O}_3 \rightarrow \text{O}_2 + \text{O}$ fast
 $\text{O} + \text{O}_3 \rightarrow 2\text{O}_2$ slow
 a) write the overall balanced chemical equation
 b) identify any intermediates within the mechanism

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 $\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$
 the reaction between Mg and HCl is represented by the equation above in a kinetics experiment
 a) 0.080 g sample of Mg ribbon is added to an erlenmeyer flask containing 50 mL of 3.0 M HCl at 25 °C

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 reactions conducted in an industrial facility mix compounds together heating and stirring them for a