

# NISSAN Learning Academy

ZD30DDTi CRD Engine

Nissan Australia July 2008

## ZD30 ENGINE CODE DESCRIPTION

### Y61 Diesel Engines .../Cont'd

**ZD30DDTi** (4 cylinder diesel engine)

**ZD** = Engine series code. The 'D' is common across all diesel engine codes

**30** = Engine capacity (3.0L)

**D** = Double Overhead camshafts

**i** = Direct Injection (as opposed to "indirect" injection as it is in the TD & RD engine series)

**T** = Turbo

**I** = Intercooler

**Note 1;** None of the coding makes any reference to Electronic Control even though it is an electronically controlled engine.

**Note 2;** The same code applies to the CRD (Common Rail Diesel) version of this engine.

**Note 3;** All ZD30 engines are fitted with NATS (wagon only).

**Note 4;** Only ZD30 CRD engines from 2007 production have ASCD (Cruise Control – wagon only).

**Note 5;** ZD30 engine also applied to D22 Navara models from late 2001 ~ 2006 production. Engine has a conventional turbo, but no intercooler. (ZD30DDT)



ZD30DDTi (Non CRD; 2000 – 2006)



ZD30DDTi (CRD; 2007)



Rocker Cover ZD30DDTi CRD (2007)



Rocker Cover ZD30DDTi (2000 – 2006)

# Nissan Zd30 Manual

**DP Hallahan**



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neutralization reaction

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web the balanced chemical equation for this reaction is  $\text{mg} + 2\text{hno}_3 \text{ aq} \rightarrow \text{mg}(\text{no}_3)_2 \text{ aq} + \text{h}_2 \text{ g}$  in this equation the s symbolizes solid magnesium aq represents aqueous nitric acid aq signifies magnesium nitrate in a dissolved state and g represents the gaseous hydrogen produced

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web magnesium nitric acid magnesium nitrate dihydrogen  $\text{mg} + \text{hno}_3 \rightarrow \text{mg}(\text{no}_3)_2 + \text{h}_2$  is a single displacement substitution reaction where one mole of solid magnesium mg and two moles of aqueous nitric acid  $\text{hno}_3$  react to form one mole of aqueous magnesium nitrate  $\text{mg}(\text{no}_3)_2$  and one mole of dihydrogen  $\text{h}_2$  gas show chemical

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web mar 4 2018  $\text{mg} + 2\text{hno}_3 \rightarrow \text{mg}(\text{no}_3)_2 + \text{h}_2$  on the left side there are mg 1 magnesium in elemental form  $\text{h}_2$  there are 2 molecules of nitric acid each acid has one hydrogen so two molecules of nitric acid is 2 hydrogens n 2 there are 2 molecules of nitric acid each acid has one nitrogen so two molecules of nitric acid is 2 nitrogens o 6

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web magnesium carbonate nitric acid magnesium nitrate water carbon dioxide one mole of magnesium carbonate  $\text{mgco}_3$  and two moles of nitric acid  $\text{hno}_3$  react to form one mole of magnesium nitrate  $\text{mg}(\text{no}_3)_2$  one mole of water  $\text{h}_2\text{o}$  and one

mole of carbon dioxide  $\text{CO}_2$  show chemical structure image reactants

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web magnesium is already balanced on the reactant and product side balancing nitrogen by multiplying  $\text{HNO}_3$  by 2 we get  
 $\text{Mg} + 2\text{HNO}_3 \rightarrow \text{Mg(NO}_3)_2 + \text{H}_2$  magnesium nitrate hydrogen gas oxygen and hydrogen get  
 automatically balanced by balancing nitrogen

*$\text{Mg(NO}_3)_2 + \text{H}_2\text{SO}_4 \rightarrow \text{MgSO}_4 + \text{HNO}_3$  chemical equation* - Apr 07 2023

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 $\text{MgCO}_3 + 2\text{H}^+ \rightarrow \text{Mg}^{2+} + \text{CO}_2 + \text{H}_2\text{O}$  write the balanced equation when nitric acid reacts with calcium hydroxide to  
 produce water and calcium nitrate

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 you balance magnesium and nitric acid you ll need to count all of atoms on both directions of the chemical equation to  
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 chemical reaction for this reaction we have a single displacement reaction balancing strategies in this single displacement  
 reaction the mg replaces the h in the  $\text{HNO}_3$

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 each compound to balance  $\text{Mg} + \text{HNO}_3 \rightarrow \text{Mg(NO}_3)_2 + \text{H}_2$  you ll need to be

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 $\text{Mg}^{2+} + 2\text{OH}^- + 2\text{H}^+ + 2\text{NO}_3^- \rightarrow \text{Mg}^{2+} + 2\text{NO}_3^- + 2\text{H}_2\text{O}$  net ionic equation  $2\text{OH}^- + 2\text{H}^+ \rightarrow 2\text{H}_2\text{O}$  sulfuric acid and lithium hydroxide  
 balanced total equation  $\text{H}_2\text{SO}_4 + 2\text{LiOH} \rightarrow \text{Li}_2\text{SO}_4 + 2\text{H}_2\text{O}$

