

RAILWAY RECRUITMENT BOARD, MALDA
LIST OF CANDIDATES IN ORDER OF NAME (ALPHABETICALLY) FOR CAT-02 (Extension Educator)
EN 23-1/2007
EXAM DATE 30/08/2009

SL. NO	CNTL	NAME	FNAME	DOB	STATUS
1	02000005	A.F. GULENUR ISLAM BARBHUIYA	FAIZUR ROHMAN BARBHUIYA	30/09/1983	ELIGIBLE
2	02000075	ALOK KUMAR TRIPATHI	PREM CHANDRA TRIPATHI	01/03/1977	ELIGIBLE
3	02000077	ALOK PRAKASH	ANIRUDH BHAGAT	01/03/1979	INELIGIBLE
4	02000032	ALOKESH BHATTACHARYYA	DHARMADAS BHATTACHARYA	09/09/1964	INELIGIBLE
5	02000023	AMIT KESARWANI	HARISH CHANDRA KESARWANI	17/07/1980	ELIGIBLE
6	02000024	AMIT KUMAR PAL	LATE PROBASH PAL	07/08/1978	ELIGIBLE
7	02000038	AMIT PANDIT	GANPATRAO PANDIT	01/04/1976	ELIGIBLE
8	02000096	AMRIT LAL MEENA	RAMJI LAL MEENA	25/01/1973	ELIGIBLE
9	02000074	ANAMIKA SINGHA ROY	SUKUMAR SINGHA ROY	28/06/1980	ELIGIBLE
10	02000053	ANANTA KUMAR SWAIN	BASANTA KUMAR SWAIN	01/06/1976	ELIGIBLE
11	02000068	ANEETA MISHRA	ALEKHA CHANDRA MISHRA	22/12/1978	ELIGIBLE
12	02000121	ANIL KUMAR	ANURUP RAM NIRANKARI	02/10/1972	ELIGIBLE
13	02000049	ANIL KUMAR SAHU	SHANKAR LAL SAHU	03/10/1973	ELIGIBLE
14	02000065	ANNAPURNA DAS	SUKUMAR DAS	21/10/1977	ELIGIBLE
15	02000041	ANNU SHRI KUMARI	RAMDUTTA SINGH	25/05/1978	INELIGIBLE
16	02000051	ANUKUL CHANDRA	RAMDEO PRASAD	22/01/1979	ELIGIBLE
17	02000043	ARGHYA SUKUL	LATE RAMRANJAN SUKUL	08/09/1978	ELIGIBLE
18	02000124	ARINDAM DAS	ASIT KUMAR DAS	29/06/1977	ELIGIBLE
19	02000021	ARUN KUMAR SHRIWASTAWA	CHANDRADEO LAL SHRIWASTAWA	01/12/1978	ELIGIBLE
20	02000031	ASHIS KUMAR GHOSH	LATE BENOY KUMAR GHOSH	16/01/1969	INELIGIBLE
21	02000128	ASHOK KUMAR	SHUBHKARAN	10/02/1979	ELIGIBLE
22	02000079	ASHOK KUMAR CHOUBEY	GANGA DHAR CHOUBEY	05/03/1975	INELIGIBLE
23	02000117	BALWAN KUMAR	AMAR NATH	12/05/1976	ELIGIBLE
24	02000070	BAPPADITYA CHATTERJEE	PROVASH CHATTERJEE	08/08/1978	ELIGIBLE
25	02000100	BHARAT SHAMRAOJI AGLAWE	SHAMRAOJI GANPAT AGLAWE	14/10/1975	ELIGIBLE
26	02000120	BHOLA CHOUDHARY	LATE GULAB CHAND CHOUDHARY	31/01/1978	ELIGIBLE
27	02000064	BINOD KUMAR SINGH	RAM KUMAR SINGH	14/01/1978	ELIGIBLE
28	02000078	BIRESH KUMAR BAIRAGI	BIMAL KRISHNA BAIRAGI	28/12/1976	ELIGIBLE
29	02000018	BRAJENDRA KUMAR	SRIGOVIND	10/06/1976	ELIGIBLE
30	02000098	CHANDANA CHOWDHURY	LATE DEBENDRANATH CHOWDHURY	21/12/1975	INELIGIBLE
31	02000052	CHANDRA PRAKASH	BHARAT LAL SAH	01/03/1977	ELIGIBLE
32	02000099	CHANDRAMANI SHYAMSUNDAR WASNIK	SHYAMSUNDAR DAJIBA WASNIK	16/08/1976	ELIGIBLE
33	02000080	DEBASMITA CHOUDHURY	PRASANTA KUMAR CHOUDHURY	06/04/1982	ELIGIBLE
34	02000112	DEEPAK KUMAR	RAMANANDI	07/08/1988	INELIGIBLE
35	02000027	DEEPAK KUMAR SINGH	RAM NARESH SINGH	25/01/1976	ELIGIBLE
36	02000029	DHRUV NATH TIWARI	RAGHAW PRASAD TIWARI	28/02/1973	ELIGIBLE

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paiditallidirector vijay gowtham rajuproducer vv

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for the nitration 4 01 ml of nitric acid 4 00 ml of sulfuric acid and 4 56 ml bromobenzene were used to complete the reaction
with these measurements bromobenzene was found to be the limiting reagent since only 4 56 ml of bromobenzene was used
the amount of product predicted to be recovered was 8 7533 g

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the aim of the experiment is to determine and study electrophilic aromatic substitution reaction eas particularly nitration of
methyl benzoate go give methyl m nitrobenzoate thus determine the weight melting point and percentage yield of the pure
methyl m nitrobenzoate abstract

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in this experiment we will carry out the nitration of bromobenzene bromine substituted on an aromatic ring is an ortho para
directing group it will direct an incoming electrophile to the ortho and para positions it is also deactivating meaning that its
reactions will be slower than the identical reactions of benzene

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section 20711 lab experiment nitration of bromobenzene introduction the chemical reaction of nitration consists of a nitro
group being added to or substituted in a molecule nitration can basically be carried out by a mixture of concentrated nitric

acid and sulfuric acid this mixture is useful to obtain the active nitronium ion

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1 go to the restroom and wash the affected area with lots of soap and water the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area 3 the summary to this warning is not to touch any part of your body while performing this experiment if you must leave the lab for any reason 1

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in the nitration of bromobenzene a mixture of sulfuric acid and nitric acid react with one another nitric acid the weaker of the two acids will serve as base to form nitronium ion bromobenzene will attack the nitronium ion by nucleophilic attack

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nitration of bromobenzene lab report improved essays 731 words 3 pages open document essay sample check writing quality show more this experiment was performed in order to demonstrate the principles of nitration through electrophilic aromatic substitution it also examined the effects of bromine on an aromatic ring ortho para vs meso directing

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then actual is your actual findings yield actual theoretical $\times 100$ what are the possible side reactions of nitration of bromobenzene dinitration can occur at ortho para or at both ortho positions occurs in overheating meta product happens very rarely because bromine is an ortho para director

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the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area theory and explanation of the nitration of bromobenzene if the temperature of the reaction mixture exceeds 60 o

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the bromine substituent in bromobenzene is an ortho para director the para position is favored due to steric effects in the reaction that takes place in this lab a nitro group is substituted onto bromobenzene at the para position to form 1 bromo 4 nitrobenzene

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study with quizlet and memorize flashcards containing terms like reacting sulfuric and nitric acid together produces the nitronium ion alone nitric acid is not a strong enough electrophile to react with bromobenzene sulfuric acid is added to the nitric acid the reverse rxn can overheat and boil easily includes a claisen adaptor and

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over 5 minutes slowly pipet this bromobenzene into the icy cold nitration mixture hno 3 h 2 so 4 in the 125 ml flask and stir continue to stir the reaction mixture in the ice bath for 30 minutes do not heat in a new 250 ml erlenmeyer

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on the right the arenium cation has a nitro set with incredibly small 0.6 angstrom bond distances this will result in an optimization failure as the atoms are even close for the programme to accomodate their electrons demonstration of creating the p arenium cation by to nitration of bromobenzene

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lab report 15 nitration of bromobenzene brynn abbott amp katie muerer bea0006 uah ch 336 02 september 21 2022 abstract in this experiment the electrophilic aromatic substitution mechanism was used to obtain 4 nitrobromobenzene and 2 nitrobromobenzene recrystallization was used to purify 4 nitrobromobenzene

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the purpose of this lab was to observe where substituents are added on a benzene ring by way of electrophilic aromatic substitution the electrophilic source in the nitration of bromobenzene comes from a nitronium ion NO_2^+ which acts as the electrophilic species

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nitration of bromobenzene is an electrophilic aromatic substitution reaction the electrophile will be formed from 4.0 ml of nitric acid and 4.0 ml of sulfuric acid and will then react with 4.5 ml of bromobenzene this will produce o bromonitrobenzene and p bromonitrobenzene

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experimental spectra reaction a fluorobenzene stock fluorobenzene 1 h 1 h 19 f 19 f 13 c nmr fid for reference only sample 1 h 1 h 19 f 19 f 13 c nmr eas nitration of fluorobenzene crude product not available for submission for credit

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chem 2261 lab 5 nitration of bromobenzene introduction an electrophilic aromatic substitution reaction is an important process that allows for intermediate chemical structures to be used as precursors to pharmaceuticals and other industrially produced chemicals by adding different functional groups into an aromatic ring

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lab report 10 nitration of bromobenzene raekwon filmore cm 244 section 40 march 27 2018 introduction for this experiment nitration of bromobenzene was the focus of the lab the benzene is an aromatic compound and when it reacts with wither a mixture of sulfuric acid or nitric acid creates what is known as a nitro group

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