

Radioactive Decay Law

When an individual nucleus transforms into another with the emission of radiation, the nucleus is said to **decay**. Radioactive decay occurs for all nuclei with $Z > 82$, and also for some unstable isotopes with $Z < 83$. The decay rate is proportional to the number of original (undecayed) nuclei N in a substance. The number of nuclei lost to decay, $-dN$ in time interval dt , is written

$$-dN = \lambda N dt$$

where λ is called the **decay constant**. (The minus sign indicates the number of original nuclei decreases over time.) In other words, the more nuclei available to decay, the more that do decay (in time dt). This equation can be rewritten as

$$\frac{dN}{N} = -\lambda dt. \quad (1)$$

Integrating both sides of the equation, and defining N_0 to be the number of nuclei at $t = 0$, we obtain

$$\int_{N_0}^N \frac{dN}{N} = -\int_0^t \lambda dt'. \quad (2)$$

This gives us

$$\ln \frac{N}{N_0} = -\lambda t. \quad (3)$$

Taking the left and right sides of the equation as a power of e , we have the **radioactive decay law**. The total number N of radioactive nuclei remaining after time t is

$$N = N_0 e^{-\lambda t} \quad (4)$$

where λ is the decay constant for the particular nucleus.

The total number of nuclei drops very rapidly at first, and then more slowly (**Figure 2**).

Radioactive Decay Law Answer Key

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