

Activity. Radioactive Decay: A Sweet Simulation of Half-Life

Purpose

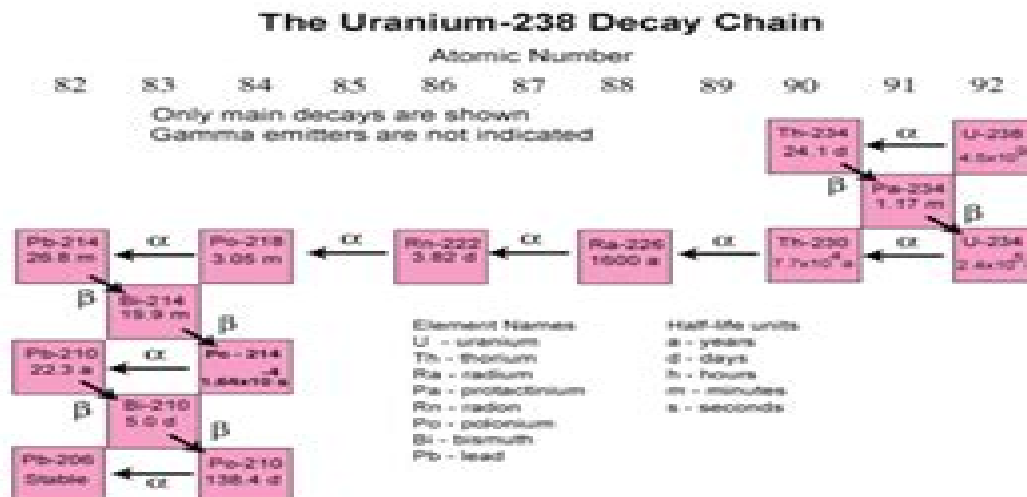
To demonstrate that the rates of decay of unstable nuclei can be measured, that the exact time that a certain nucleus will decay cannot be predicted, and that it takes a very large number of nuclei to find the rate of decay.

Background

Radioactivity is the spontaneous decay of a nucleus to form another nucleus and a nuclear particle. Some nuclides are radioactive because the ratio of the number from a minimum of 1:1 to a maximum of about 1.6 :1. All nuclides with more than 83 protons are radioactive meaning that all isotopes of the elements beyond bismuth are radioactive. These nuclei will spontaneously decay no matter how many neutrons are present.

In a collection of radioactive nuclei, the nuclei decay by splitting off nuclear particles and forming new nuclei. A sample of radioactive material will continue to decay over time until all the nuclei have decayed. For some nuclides, this will happen within seconds, minutes, hours or days. For others it may require months, years or thousands of years. Half-life is the time required for the decay of one-half of a sample of a radioactive substance. For example, if you have a 10.0 g sample of radioactive material with a half-life of four days, at the end of four days, the amount of the sample has decreased by one-half to 5.0 g. In another four days, the amount remaining is 2.5 g, etc. until all of the radioactive material has decayed or formed a more stable nucleus.

For example, uranium-238 will undergo radioactive decay as illustrated below until it forms the stable atom Pb-206.



As shown in the figure above, in a collection of radioactive nuclei, the nuclei decay by splitting off nuclear particles (α , β or γ) and forming new nuclei. The sample of radioactive material will continue to decay over time until all the nuclei have

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Answer Key

KJ Lindholm-Leary



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