

- 1.1 Find the rate of reaction defined as

$$r_{O_2} = \frac{\text{mol } O_2 \text{ used}}{\text{sec} \cdot \text{m}^3 \text{ of tank}}$$

Evaluate terms

$$\bar{t} = \frac{V}{\bar{v}} \quad \text{or} \quad V = \bar{t} \bar{v}$$

or $\left(\text{Volume of treatment tanks} \right) = \left(\frac{1}{3} \text{ day} \right) \left(32000 \frac{\text{m}^3}{\text{day}} \right) = 10667 \text{ m}^3$

O_2 used $\left(200 \frac{\text{mg}}{\text{lit}} \right) \left(\frac{1 \text{ gm}}{1000 \text{ mg}} \right) \left(\frac{\text{mol}}{32 \text{ gm}} \right) \left(\frac{1000 \text{ lit}}{\text{m}^3} \right) \left(\frac{32000 \text{ m}^3}{\text{day}} \right) = 2 \times 10^5 \frac{\text{mol } O_2}{\text{day}}$

Thus the rate of reaction

$$\left. \begin{aligned} \frac{2.0 \times 10^5 \text{ mol } O_2 / \text{day}}{10667 \text{ m}^3} &= 18.75 \text{ mol/m}^3 \cdot \text{day} \\ &= 2.17 \times 10^{-4} \text{ mol/m}^3 \cdot \text{s} \end{aligned} \right\} \leftarrow$$

- 1.3 Find $-r'_{C_{20}H_{42}}$ and $-r^M_{C_{20}H_{42}}$... evaluate terms

$$V_{\text{cat}} = \frac{50000 \text{ kg}}{800 \text{ kg/m}^3} = 62.5 \text{ m}^3 \text{ of catalyst}$$

$$W_{\text{cat}} = 50000 \text{ kg}$$

$$\text{MW}_{C_{20}H_{42}} = [20(12) + 42(1)] \frac{1}{1000} = 0.282 \text{ kg/mol}$$

$$F_{\text{feed}} = (6000 \text{ m}^3/\text{day})(900 \text{ kg/m}^3) = 5400000 \text{ kg/day}$$

So $-\frac{dN_{C_{20}H_{42}}}{dt} = \left(\frac{5400000 \text{ kg/day}}{0.282 \text{ kg/day}} \right) \left(\frac{\text{day}}{24(3600) \text{ s}} \right) (0.6) = 133 \text{ mol reacted/s}$

Thus the rate of disappearance of $C_{20}H_{42}$

$$\left. \begin{aligned} -r' &= \frac{1}{W_{\text{cat}}} \frac{dN}{dt} = \frac{1}{50000} (133) = 0.0027 \text{ mol/kg cat} \cdot \text{s} \\ -r^M &= \frac{1}{V_{\text{cat}}} \frac{dN}{dt} = \frac{1}{62.5} (133) = 2.13 \text{ mol/m}^3 \text{ cat} \cdot \text{s} \end{aligned} \right\} \leftarrow$$

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Chemical Engineering Education, 1984 *Engineering and Chemical Thermodynamics* Milo D. Koretsky, 2012-12-17
Koretsky helps students understand and visualize thermodynamics through a qualitative discussion of the role of molecular interactions and a highly visual presentation of the material. By showing how principles of thermodynamics relate to molecular concepts learned in prior courses, *Engineering and Chemical Thermodynamics 2e* helps students construct new knowledge on a solid conceptual foundation. *Engineering and Chemical Thermodynamics 2e* is designed for Thermodynamics I and Thermodynamics II courses taught out of the Chemical Engineering department to Chemical Engineering majors. Specifically designed to accommodate students with different learning styles, this text helps establish a solid foundation in engineering and chemical thermodynamics. Clear conceptual development, worked-out examples, and numerous end-of-chapter problems promote deep learning of thermodynamics and teach students how to apply thermodynamics to real-world engineering problems.

Furfural: An Entry Point Of Lignocellulose In Biorefineries To Produce Renewable Chemicals, Polymers, And Biofuels Manuel Lopez Granados, David Martin Alonso, 2018-06-19
There is a wide consensus that furfural, a renewable commodity currently obtained from lignocellulosic agro-residues with a production volume of around 300 kTon per year, is a key feedstock for leveraging lignocellulosic residues in future biorefineries. Several chemicals are already being manufactured from furfural due to its advantageous production cost. Furthermore, a vast number of others are also technically viable to produce from oil. This book compiles the vast existing information into relevant stages of transformations of furfural as renewable chemicals, biofuels, and bioresins, focusing on the relevant chemical and engineering aspects of processes to obtain them, including reactors and catalysis. It offers essential information for improving the economic and environmental viability of current commercial applications and upcoming future applications. It should be of particular interest to graduate and advanced undergraduate students as well as engineers and academic researchers alike who are working in the field.

Nigerian Journal of Research and Production, 2007 *Aplikasi Spreadsheet untuk Perhitungan Teknik Kimia Sederhana* Rama Oktavian, A.S. Dwi Saptati Nur Hidayati, 2017-08-01
Buku ini membahas tentang aplikasi spreadsheet yang disediakan oleh Microsoft Excel 2013 untuk membantu menyelesaikan perhitungan yang terkait dengan permasalahan di bidang Teknik Kimia. Pembahasan di dalam buku ini dibagi menjadi 5 bab, yaitu pada Bab 1 dibahas mengenai aplikasi spreadsheet yang digunakan di perhitungan Teknik Kimia dan dasar. Bab 2 menjelaskan tentang prinsip-prinsip dasar perhitungan neraca massa di Teknik Kimia dan aplikasi spreadsheet untuk perhitungan neraca massa. Bab 3 berisi tentang prinsip-prinsip dasar perhitungan neraca energi di Teknik Kimia dan aplikasi spreadsheet untuk perhitungan neraca energi pada sistem tanpa reaksi dan dengan reaksi. Bab 4 menjelaskan tentang prinsip-prinsip dasar termodinamika teknik kimia, istilah termodinamika, properti termodinamika, persamaan keadaan, dan kesetimbangan fasa, serta aplikasi spreadsheet yang digunakan untuk perhitungan properti termodinamika, persamaan keadaan, dan kesetimbangan fasa. Bab 5

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