

## Relative Formula Mass

If  $H_2O$  is the chemical formula for water, how many atoms are in total in the substance? **3**

In chemical symbols each capital letter indicates the next element. Use a periodic table to complete the following. For the following substance list the names of the elements and the number of atoms present in the chemical formula.

$KNO_3$  \_\_\_\_\_

\_\_\_\_\_

\_\_\_\_\_

The relative atomic mass,  $A_r$ , is the mean mass of an atom of an element compared with one twelfth of

the mass of a \_\_\_\_\_  $^{12}C$  atom.

Circle on the Calcium tile the relative atomic mass.



Relative formula mass is the mean mass of a substance if all

the relative atomic masses are \_\_\_\_\_ together.

The relative atomic mass of carbon is 12 and oxygen is 16.

The relative formula mass of  $CO_2$  is  $12 + (2 \times 16) = 44$ .

Complete the following (Use a periodic table):

$CH_4$  \_\_\_\_\_

\_\_\_\_\_

$MgCl_2$  \_\_\_\_\_

\_\_\_\_\_

$Mg(NO_3)_2$  \_\_\_\_\_

\_\_\_\_\_

## Empirical Formula

Empirical formula shows the simplest whole number ratio of atoms of each element in a compound.

Look at the space saving models of Butane.

Number of carbon atoms \_\_\_\_\_

Number of hydrogen atoms \_\_\_\_\_

Chemical formula \_\_\_\_\_ Empirical formula \_\_\_\_\_

Look at the displayed formula of valeric acid.

Number of carbon atoms \_\_\_\_\_

Number of hydrogen atoms \_\_\_\_\_

Number of oxygen atoms \_\_\_\_\_

Chemical formula \_\_\_\_\_ Empirical formula \_\_\_\_\_

Derive the empirical formula for the following:

$Al_2O_3$  \_\_\_\_\_  $C_4H_8O_2$  \_\_\_\_\_

$Ca_3(PO_4)_2$  \_\_\_\_\_  $Na_2SO_4$  \_\_\_\_\_

## C2.1 - Separating Mixtures - Combined Chemistry Revision Sheet

### Pure and Impure Substances

Complete the sentences using the words in the box. Use each word once, more than once or not at all.

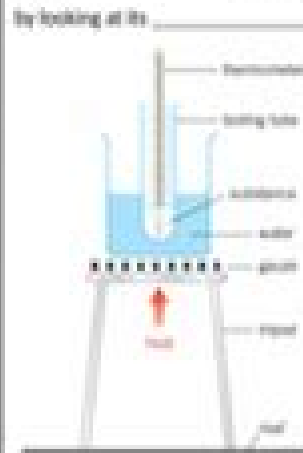
metals steel mineral mixed harder impure one melting  
many iron pure joined distilled alloys sulfur

In science a pure substance consists of \_\_\_\_\_ element/s or compound/s.

Mixtures are \_\_\_\_\_. They contain two or more different elements or compounds \_\_\_\_\_ together. \_\_\_\_\_ water and \_\_\_\_\_ are mixtures. Mixtures of metals, called \_\_\_\_\_

are \_\_\_\_\_, than pure metals. We often determine purity of a substance

by looking at its \_\_\_\_\_ point.



Three students, A, B and C make aspirin and check its purity. The students test the purity of the aspirin by measuring the melting point, using the equipment shown. Their results are shown in the table.

What two things do they need to record to draw a change of state graph?

\_\_\_\_\_

\_\_\_\_\_

What do they need to do during the experiment to improve accuracy?

\_\_\_\_\_

\_\_\_\_\_

Student	Melting point ( $^{\circ}C$ )
A	140
B	139
C	137

The melting point of pure aspirin is in the range  $138-140^{\circ}C$ . Which student made pure aspirin?

### Crystallisation

Number the following steps to show the order you would carry them out to produce a sample of pure copper sulfate crystals. The first step has been done for you.



Label the diagram below using the labels in the box once.

Bunsen burner tripod  
gauze evaporating basin

- ☐ Add an excess of copper carbonate powder to the acid and stir gently with a stirring rod.
- ☐ Carefully warm the solution of copper sulfate so that the water evaporates.
- ☐ Stop heating when there is a teaspoon of liquid left, allow the copper sulfate to cool and crystallise fully.
- 1** Place 25 cm<sup>3</sup> sulfuric acid solution into a beaker.
- ☐
- ☐ Fold a filter paper and place it into the filter funnel.

Carefully pour the mixture through the filter paper to remove excess copper carbonate powder. The filtrate passes through the filter paper and is

## Filtering

Label the diagram below using the labels in the box once.

funnel filtrate beaker  
conical flask filter paper  
solution residue



Answer the following questions about when substances dissolve.

When a substance dissolves in another substance a \_\_\_\_\_ is made. A substance which

can be dissolved is said to be \_\_\_\_\_.

The substance that dissolves is called the \_\_\_\_\_.

The substance it dissolves in is called the \_\_\_\_\_.

When a substance does not dissolve it is said to be \_\_\_\_\_.

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**Puja Mehta**

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### Table of Contents **Ocr Gateway Past Papers Chemistry**

1. Understanding the eBook **Ocr Gateway Past Papers Chemistry**
  - The Rise of Digital Reading **Ocr Gateway Past Papers Chemistry**
  - Advantages of eBooks Over Traditional Books
2. Identifying **Ocr Gateway Past Papers Chemistry**
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an **Ocr Gateway Past Papers Chemistry**
  - User-Friendly Interface
4. Exploring eBook Recommendations from **Ocr Gateway Past Papers Chemistry**
  - Personalized Recommendations
  - **Ocr Gateway Past Papers Chemistry** User Reviews and Ratings
  - **Ocr Gateway Past Papers Chemistry** and Bestseller Lists

5. Accessing Ocr Gateway Past Papers Chemistry Free and Paid eBooks
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  - Ocr Gateway Past Papers Chemistry Budget-Friendly Options
6. Navigating Ocr Gateway Past Papers Chemistry eBook Formats
  - ePub, PDF, MOBI, and More
  - Ocr Gateway Past Papers Chemistry Compatibility with Devices
  - Ocr Gateway Past Papers Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ocr Gateway Past Papers Chemistry
  - Highlighting and Note-Taking Ocr Gateway Past Papers Chemistry
  - Interactive Elements Ocr Gateway Past Papers Chemistry
8. Staying Engaged with Ocr Gateway Past Papers Chemistry
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ocr Gateway Past Papers Chemistry
9. Balancing eBooks and Physical Books Ocr Gateway Past Papers Chemistry
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ocr Gateway Past Papers Chemistry
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Ocr Gateway Past Papers Chemistry
  - Setting Reading Goals Ocr Gateway Past Papers Chemistry
  - Carving Out Dedicated Reading Time
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  - Fact-Checking eBook Content of Ocr Gateway Past Papers Chemistry
  - Distinguishing Credible Sources
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- Exploring Educational eBooks

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