

BRILLIANT FIELD SYNERGY ACADEMY
SECOND TERM EXAMINATION 2023/2024 ACADEMIC SESSION

SUBJECT: CHEMISTRY

CLASS: SS1

INSTRUCTION: ANSWER ALL QUESTIONS

1. The following are types of salts EXCEPT (A) acidic (B) basic (C) complex (D) mono salt
2. The gas law which describes the relationship between volume and temperature is (A) Boyle's Law (B) Charles's law (C) Dalton's Law (D) Graham's Law
3. Which of the following phenomena leads to decrease in volume of a liquid in an open container? (A) Brownian motion (B) Diffusion (C) Evaporation (D) Sublimation
4. The following are allotropes of carbon EXCEPT (A) Diamond (B) Graphite (C) Coke (D). Feldspar
5. In which of these substances is the basicity of the acid equal to 1
(A) H_2SO_4 (B) H_3PO_4 (C) CH_3COOH
6. The following are physical properties of acids EXCEPT
(A) acid turns blue litmus paper red (B) concentrated acids are corrosive
(C) acid bleaches substances (D) dilute acids have sour taste
7. Calculate the pressure of Nitrogen gas if its volume is doubled given that its volume is 10dm^3 at a pressure of 6atm (A) 3atm (B) 7atm (C) 5.5atm (D) 3.3atm
8. Which of these shows a mathematical statement of Boyle's law
(A) $PV = \text{constant}$ (B) $P/V = \text{constant}$ (C) $PV^2 = \text{constant}$ (D) $P/T = \text{constant}$
9. A reaction between acids and bases are frequently called (A) combustion
(B) neutralization (C) oxidation (D) substitution
solution (D) petrol
10. Calculate the pH of 0.1mol/dm^3 of solution of HCl (A) 1 (B) 2 (C) 3 (D) 1.5

PART B - ESSAY

INSTRUCTION: Answer only 5 questions

- 1 (i) state Dalton's law of partial pressure (ii) calculate the partial pressures of three gases collected at 760mmHg if the gases are made of 40% oxygen, 30% nitrogen and 30% hydrogen respectively
- 2 (i) state the equation of the general gas law and define all the terms (ii) From the general gas equation, calculate the volume of one mole of an ideal gas at s.t.p ($R = 0.082\text{dm}^3\text{atm K}^{-1}\text{mol}^{-1}$, $T = 273\text{K}$, $p = 1\text{atm}$)

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KJ Lindholm-Leary



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Table of Contents Qustion For Second Term Exam For Ss1 2015

1. Understanding the eBook Qustion For Second Term Exam For Ss1 2015
 - The Rise of Digital Reading Qustion For Second Term Exam For Ss1 2015
 - Advantages of eBooks Over Traditional Books
2. Identifying Qustion For Second Term Exam For Ss1 2015
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Qustion For Second Term Exam For Ss1 2015
 - User-Friendly Interface
4. Exploring eBook Recommendations from Qustion For Second Term Exam For Ss1 2015

- Personalized Recommendations
 - Question For Second Term Exam For Ss1 2015 User Reviews and Ratings
 - Question For Second Term Exam For Ss1 2015 and Bestseller Lists
5. Accessing Question For Second Term Exam For Ss1 2015 Free and Paid eBooks
- Question For Second Term Exam For Ss1 2015 Public Domain eBooks
 - Question For Second Term Exam For Ss1 2015 eBook Subscription Services
 - Question For Second Term Exam For Ss1 2015 Budget-Friendly Options
6. Navigating Question For Second Term Exam For Ss1 2015 eBook Formats
- ePub, PDF, MOBI, and More
 - Question For Second Term Exam For Ss1 2015 Compatibility with Devices
 - Question For Second Term Exam For Ss1 2015 Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Question For Second Term Exam For Ss1 2015
 - Highlighting and Note-Taking Question For Second Term Exam For Ss1 2015
 - Interactive Elements Question For Second Term Exam For Ss1 2015
8. Staying Engaged with Question For Second Term Exam For Ss1 2015
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Question For Second Term Exam For Ss1 2015
9. Balancing eBooks and Physical Books Question For Second Term Exam For Ss1 2015
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Question For Second Term Exam For Ss1 2015
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Question For Second Term Exam For Ss1 2015
- Setting Reading Goals Question For Second Term Exam For Ss1 2015
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Question For Second Term Exam For Ss1 2015

- Fact-Checking eBook Content of Qustion For Second Term Exam For Ss1 2015
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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