

Stoichiometry Reflect #2

Learning Target... I can use stoichiometry to calculate amounts

1a. Balance this equation



1b. Determine how many moles of ammonium oxide was used if the student weighed out 55.2 g of it.

$$55.2 \text{ g (NH}_4\text{)}_2\text{O} \left(\frac{1 \text{ mol A}}{52.10 \text{ g}} \right) = \boxed{1.06 \text{ moles (NH}_4\text{)}_2\text{O}}$$

1c. Determine the amount of moles of ammonium nitride produced in this reaction.

$$1.06 \text{ moles A} \left(\frac{2 \text{ mol B}}{3 \text{ mol A}} \right) = \boxed{0.706 \text{ moles (NH}_4\text{)}_3\text{N}}$$

1d. Determine how many grams of ammonium nitride produced in this reaction.

$$0.706 \text{ moles} \left(\frac{108.11 \text{ g}}{1 \text{ mol B}} \right) = \boxed{48.1 \text{ g (NH}_4\text{)}_3\text{N}}$$

2. Using the following chemical reaction:

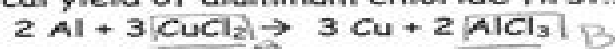


and, knowing that 61.6 g of copper was produced in this experiment, how many molecules of mercury nitrate was also produced?

$$61.6 \text{ g Cu} \left(\frac{1 \text{ mol A}}{63.55 \text{ g}} \right) \left(\frac{2 \text{ mol B}}{1 \text{ mol A}} \right) \left(\frac{6.02 \times 10^{23} \text{ mc}}{1 \text{ mol B}} \right) = \boxed{1.17 \times 10^{24} \text{ mc HgNO}_3}$$

3. Challenge...

What is the percent yield of the following reaction if 12.8 g of copper (II) chloride was used in the lab, and 7.75 g of aluminum chloride was measured at the end of the experiment? (hint: you must calculate the theoretical yield of aluminum chloride first!!)



$$12.8 \text{ g A} \left(\frac{1 \text{ mol A}}{134.45 \text{ g}} \right) \left(\frac{2 \text{ mol B}}{3 \text{ mol A}} \right) \left(\frac{133.33 \text{ g}}{1 \text{ mol B}} \right) = 8.46 \text{ g AlCl}_3$$

theoretical

$$\% \text{ yield} = \frac{7.75 \text{ g}}{8.46 \text{ g}} = \boxed{91.4\%}$$

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Heather Hattori, Richard H. Langley



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and learning Some of these include helpful explanations many examples alternative ways to solve problems plenty of practice questions complete answers an appendices With this book you will be well prepared for your exams and boost your performance

CONTENTS 1 Writing the Formulae of Ionic Compounds 2 Formula Masses 3 Mass and Percent Compositions 4 The Mole and Mass 5 The Mole and Number of Particles 6 The Mole and Concentration 7 The Mole and Molar Volume 8 Empirical and Molecular Formulae 9 Chemical Equations 10 Calculations Based on Chemical Equations 11 Limiting Reagent 12 Gas Volume Volume Stoichiometry 13 Percent Yield Answers to Practice Problems Appendices

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Table of Contents Practice Problems Stoichiometry Glencoe Mcgraw Hill Answers

1. Understanding the eBook Practice Problems Stoichiometry Glencoe Mcgraw Hill Answers
 - The Rise of Digital Reading Practice Problems Stoichiometry Glencoe Mcgraw Hill Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Practice Problems Stoichiometry Glencoe Mcgraw Hill Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Practice Problems Stoichiometry Glencoe Mcgraw Hill Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Practice Problems Stoichiometry Glencoe Mcgraw Hill Answers

- Personalized Recommendations
 - Practice Problems Stoichiometry Glencoe Mcgraw Hill Answers User Reviews and Ratings
 - Practice Problems Stoichiometry Glencoe Mcgraw Hill Answers and Bestseller Lists
5. Accessing Practice Problems Stoichiometry Glencoe Mcgraw Hill Answers Free and Paid eBooks
 - Practice Problems Stoichiometry Glencoe Mcgraw Hill Answers Public Domain eBooks
 - Practice Problems Stoichiometry Glencoe Mcgraw Hill Answers eBook Subscription Services
 - Practice Problems Stoichiometry Glencoe Mcgraw Hill Answers Budget-Friendly Options
 6. Navigating Practice Problems Stoichiometry Glencoe Mcgraw Hill Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Practice Problems Stoichiometry Glencoe Mcgraw Hill Answers Compatibility with Devices
 - Practice Problems Stoichiometry Glencoe Mcgraw Hill Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Practice Problems Stoichiometry Glencoe Mcgraw Hill Answers
 - Highlighting and Note-Taking Practice Problems Stoichiometry Glencoe Mcgraw Hill Answers
 - Interactive Elements Practice Problems Stoichiometry Glencoe Mcgraw Hill Answers
 8. Staying Engaged with Practice Problems Stoichiometry Glencoe Mcgraw Hill Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Practice Problems Stoichiometry Glencoe Mcgraw Hill Answers
 9. Balancing eBooks and Physical Books Practice Problems Stoichiometry Glencoe Mcgraw Hill Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Practice Problems Stoichiometry Glencoe Mcgraw Hill Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Practice Problems Stoichiometry Glencoe Mcgraw Hill Answers
 - Setting Reading Goals Practice Problems Stoichiometry Glencoe Mcgraw Hill Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Practice Problems Stoichiometry Glencoe Mcgraw Hill Answers

- Fact-Checking eBook Content of Practice Problems Stoichiometry Glencoe Mcgraw Hill Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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