



COVALENT BONDING

Practice Problems

Do your answers to these very carefully.

SECTION 8.1 MOLECULAR COMPOUNDS

1. Classify each of the following as ionic or molecular.

- a. NaCl (ionic) b. H_2O (molecular) c. HCl (molecular)
 d. KNO_3 (ionic) e. H_2SO_4 (molecular)

2. Which of the following is a molecular compound?

- a. NaCl b. H_2O c. HCl
 d. KNO_3 e. H_2SO_4

3. Which of the following is a molecular compound? Write the name of the compound if molecular.
 a. H_2O (molecular) b. H_2SO_4 (molecular) c. HCl (molecular) d. KNO_3 (ionic)
 e. NaCl (ionic) f. H_2SO_4 (molecular) g. HCl (molecular) h. KNO_3 (ionic)

SECTION 8.2 THE NATURE OF COVALENT BONDING

1. Draw the Lewis structure of hydrogen fluoride, HF .



2. Draw the Lewis structure of methane, CH_4 .



3. Draw the Lewis structure of carbon dioxide, CO_2 .



4. Draw the Lewis structure of acetylene, C_2H_2 .

5. Draw the Lewis structure of ethane, C_2H_6 . Do you see any structural isomers? Draw.



(None are isomers.)

SECTION 8.3 BONDING: THE ORBITAL

1. Which orbital has two regions of electron density, one above and one below the nucleus?
 a. $1s$ b. $2s$ c. $2p$ d. $3p$ e. $4p$

Practice Problems Chapter 8 Covalent Bonding Worksheet

M Woodhall



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