Thermochemistry Review Questions

- The process of dissolving ammonium nitrate, NH₄NO₃, in water is an <u>endothermic</u> process. What is the sign of 'q'? If you were to add some ammonium nitrate to water in a flask, would you expect the flask to feel warm or cold?
- Nitric acid, a source of many nitrogen compounds, is produced from nitrogen dioxide. An old process for
 making nitrogen dioxide employed nitrogen and oxygen: N_{2(g)} + 2O_{2(g)} → 2NO_{2(g)}. The reaction absorbs 56.2
 kJ of heat per 2 mol of NO_{2(g)} produced. Is the reaction endothermic or exothermic? What is the value of q?
- Ammonia burns in the presence of a copper catalyst to form nitrogen gas: NH_{3(g)} + 3O_{2(g)} → 2N_{2(g)} + 6H₂O_(g) with a ΔH = −1267 kJ/mol. What is the enthalpy change (ΔH) to burn 35.8 g of ammonia? [Ans: -2 663.45 kJ]
- Propane, C₃H₈, is a common fuel gas. Use the following to calculate the grams of propane you would need to provide 369 kJ of heat: C₃H_{8(a)} + 5O_{2(a)} → 3CO_{2(a)} + 4H₂O_(a) with a ΔH = −2043 kJ/mol. [Ans: 7.96 g]
- Ammonia burns in the presence of a platinum catalyst to form nitric oxide: 4NH_{3(g)} + 5O_{2(g)} → 4NO_(g) + 6H₂O_(g). What is the heat of reaction at constant pressure? Use the following thermochemical equations [Ans: -906.3 kJ]:

$$N_{2(g)} + O_{2(g)} \rightarrow 2NO_{(g)}$$
 $\Delta H = 180.6 \ kJ$
 $N_{2(g)} + 3H_{2(g)} \rightarrow 2NH_{3(g)}$ $\Delta H = -91.8 \ kJ$
 $2H_{2(g)} + O_{2(g)} \rightarrow 2H_{2}O_{(g)}$ $\Delta H = -483.7 \ kJ$

Hydrogen sulfide, H₂S, is a poisonous gas with the odour of rotten eggs. The reaction for the formation of H₂S from the element is: H_{2(g)} + ½S_{B(rhombic)} → H₂S_(g). Use Hess' law to obtain the enthalpy change for this reaction from the following enthalpies [Ans: - 21 kJ]:

$$\begin{array}{ll} H_2S_{(g)} + \frac{3}{2}O_{2(g)} \to H_2O_{(g)} + SO_{2(g)} & \Delta H = -518 \, kJ \\ H_{2(g)} + \frac{1}{2}O_{2(g)} \to H_2O_{(g)} & \Delta H = -242 \, kJ \\ \frac{1}{8}S_{8(rhomble)} + O_{2(g)} \to SO_{2(g)} & \Delta H = -297 \, kJ \end{array}$$

- The cooling effect of alcohol on the skin is due to its evaporation. Calculate the heat of vaporization (ΔH) of ethanol, C₂H₅OH_(I) → C₂H₅OH_(g). The standard enthalpy of C₂H₅OH_(I) is 277.7 kJ/mol and that of C₂H₅OH_(g) is 235.4 kJ/mol. [Ans: 42.3 kJ]
- Carbon disulfide is a colourless liquid. When pure, it is nearly odourless, but the commercial produce smells vile. Carbon disulfide is used in the manufacturer of rayon and cellophane. The liquid burns as follows: CS_{2(I)} + 3O_{2(g)} → CO_{2(g)} + 2SO_{2(g)}. Calculate the standard enthalpy change for this reaction using standard enthalpies of formation [Ans: -1076.6 kJ].
- The group IIA carbonates decompose when heated. For example: BaCO_{3(s)} → BaO_(s) + CO_{2(g)}. Use the following enthalpies of formation and calculate the heat require to decompose 10 g of barium carbonate. The standard enthalpy of BaCO₃ is 1216.3 kJ/mol. BaO_(s) is 942.3 kJ/mol and CO_{2(g)} is 393.5 kJ/mol. [Ans: -6.09 kJ].

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