

- (d) Write the chemical equation of the reaction which takes place on reacting XNO_3 solution and common salt solution giving the physical states of all the reactants and products.
- (e) What type of chemical reaction is illustrated by the above equation?
64. Two metals X and Y form the salts XSO_4 and Y_2SO_4 respectively. The solution of salt XSO_4 is blue in colour whereas that of Y_2SO_4 is colourless. When barium chloride solution is added to XSO_4 solution, then a white precipitate Z is formed along with a salt which turns the solution green. And when barium chloride solution is added to Y_2SO_4 solution, then the same white precipitate Z is formed along with colourless common salt solution.
- (a) What could the metals X and Y be?
- (b) Write the name and formula of salt XSO_4 .
- (c) Write the name and formula of salt Y_2SO_4 .
- (d) What is the name and formula of white precipitate Z?
- (e) Write the name and formula of the salt which turns the solution green in the first case.
65. A red-brown metal X forms a salt XSO_4 . When hydrogen sulphide gas is passed through an aqueous solution of XSO_4 , then a black precipitate of XS is formed along with sulphuric acid solution.
- (a) What could the salt XSO_4 be?
- (b) What is the colour of salt XSO_4 ?
- (c) Name the black precipitate XS .
- (d) By using the formula of the salt obtained in (a) above, write an equation of the reaction which takes place when hydrogen sulphide gas is passed through its aqueous solution.
- (e) What type of chemical reaction takes place in this case?
66. When a strip of red-brown metal X is placed in a colourless salt solution YNO_3 , then metal Y is set free and a blue coloured salt solution $\text{X(NO}_3)_2$ is formed. The liberated metal Y forms a shining white deposit on the strip of metal X.
- (a) What do you think metal X is?
- (b) Name the salt YNO_3 .
- (c) What could be metal Y?
- (d) Name the salt $\text{X(NO}_3)_2$.
- (e) What type of reaction takes place between metal X and salt solution YNO_3 ?
67. A metal salt MX when exposed to light splits up to form metal M and a gas X_2 . Metal M is used in making ornaments whereas gas X_2 is used in making bleaching powder. The salt MX is itself used in black and white photography.
- (a) What do you think metal M is?
- (b) What could be gas X_2 ?
- (c) Name the metal salt MX .
- (d) Name any two salt solutions which on mixing together can produce a precipitate of salt MX .
- (e) What type of chemical reaction takes place when salt MX is exposed to light? Write the equation of the reaction.

ANSWERS

1. Decomposition reaction 4. $\text{Zn} + 2\text{AgNO}_3 \longrightarrow \text{Zn(NO}_3)_2 + 2\text{Ag}$ 5. Rancidity 6. Anti-oxidants
11. Decomposition reactions (carried out by electrolysis) 14. $2\text{AgNO}_3(\text{aq}) + \text{Cu}(\text{s}) \longrightarrow \text{Cu(NO}_3)_2(\text{aq}) + 2\text{Ag}(\text{s})$ 15. (i) Decomposition (ii) Combination (iii) Decomposition (iv) Decomposition (v) Decomposition (vi) Decomposition
- (c) Combination 16. (a) Combination (b) Decomposition (c) Decomposition (d) Decomposition
- (e) Combination 17. (i) Displacement reaction (ii) Combination reaction (iii) Decomposition reaction
- reaction (iv) Double displacement reaction (v) Displacement reaction 18. (a) See page 29 (b) See page 29
19. (a) Displacement reaction (b) Combination reaction 20. (a) Combination reaction (b) Displacement reaction (c) Displacement reaction (d) Decomposition reaction (e) Double displacement reaction
21. (a) H_2O_2 (b) PbS 22. H_2S 23. Substance oxidised : H_2S ; Substance reduced : SO_2
24. (a) oxidation ; reduction (b) reduction ; oxidation (c) rancidity 25. (i) C (ii) ZnO 26. (b) Magnesium is oxidised ; Addition of oxygen to magnesium takes place (c) (i) HCl (ii) MnO_2 (iii) MnO_2 (iv) HCl
31. $\text{ZnO} + \text{C} \longrightarrow \text{Zn} + \text{CO}$ (i) Zinc oxide (ii) Carbon 34. (b) $2\text{Cu} + \text{O}_2 \longrightarrow 2\text{CuO}$ (i) Cu (ii) O_2
35. (b) $2\text{Mg} + \text{O}_2 \longrightarrow 2\text{MgO}$ (i) Mg (ii) O_2 37. (b) $\text{Cu}(\text{s}) + 2\text{AgNO}_3(\text{aq}) \longrightarrow \text{Cu(NO}_3)_2(\text{aq}) + 2\text{Ag}(\text{s})$ 39. (a) H_2 (b) CuO (c) CuO (d) H_2 43. (b) Rusting (c) Oxidation 44. (b) Oxidation

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Table of Contents S Chand Lab Manual Class 5 Science

1. Understanding the eBook S Chand Lab Manual Class 5 Science
 - The Rise of Digital Reading S Chand Lab Manual Class 5 Science
 - Advantages of eBooks Over Traditional Books
2. Identifying S Chand Lab Manual Class 5 Science
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an S Chand Lab Manual Class 5 Science
 - User-Friendly Interface
4. Exploring eBook Recommendations from S Chand Lab Manual Class 5 Science
 - Personalized Recommendations
 - S Chand Lab Manual Class 5 Science User Reviews and Ratings
 - S Chand Lab Manual Class 5 Science and Bestseller Lists
5. Accessing S Chand Lab Manual Class 5 Science Free and Paid eBooks

- S Chand Lab Manual Class 5 Science Public Domain eBooks
- S Chand Lab Manual Class 5 Science eBook Subscription Services
- S Chand Lab Manual Class 5 Science Budget-Friendly Options
- 6. Navigating S Chand Lab Manual Class 5 Science eBook Formats
 - ePub, PDF, MOBI, and More
 - S Chand Lab Manual Class 5 Science Compatibility with Devices
 - S Chand Lab Manual Class 5 Science Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of S Chand Lab Manual Class 5 Science
 - Highlighting and Note-Taking S Chand Lab Manual Class 5 Science
 - Interactive Elements S Chand Lab Manual Class 5 Science
- 8. Staying Engaged with S Chand Lab Manual Class 5 Science
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers S Chand Lab Manual Class 5 Science
- 9. Balancing eBooks and Physical Books S Chand Lab Manual Class 5 Science
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection S Chand Lab Manual Class 5 Science
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine S Chand Lab Manual Class 5 Science
 - Setting Reading Goals S Chand Lab Manual Class 5 Science
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of S Chand Lab Manual Class 5 Science
 - Fact-Checking eBook Content of S Chand Lab Manual Class 5 Science
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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