Model 2 – Ternary Acids (Oxyacids)

Acid	Name of Acid in Aqueous Solution	Cation (+)	Polyatomic Anion	Polyatomic Anion Name
HCIO.	Chloric acid	H,O	C/O3-1	Chlorate
H.SO.	Sulfurous acid	24,00	503-2	Sulfise
H.SO.	Sulfuric acid	2 Hso*	504-3	Sulfate
H,PO,	Phosphorous acid	3 H,O	P03-3	Phosphite
H.PO.	Phosphoric acid	3 H50+	PO _a s-	Phosphate
HNO,	Nitric acid	H30+	NO."	Nitrate
HNO,	Nitrous acid	H.0*	NO,	Nitrite
H,CO,	Carbonic acid	2450+	CO,-2	Carbonate

- 6. Look at the formulas of the ternary acids in Model 2.
 - w. How are ternary acids different from binary acids in their structure?

Ternory acids contain three elements and they also contain the element oxygen

6. What number do you think the prefix "ter-" refers to?

"ter" rofers to three

- 7. When ternary acids are mixed with water, ions will form. Fill in the table above with the formulas and names of the anions.
- 8. Examine the pairs of ternary acids in Model 2 that contain sulfur, phosphorus, and nitrogen. Each pair has one acid that ends in "-ic" and another that ends in "-ous." These endings are related to the name of the polyatomic anion found in the acid ("-ate" or "-ite"). Complete the statements below with the correct acid name ending.

Polyatomic anion ending is "-ate" → acid name ending is /C

Polyatomic anion ending is "-ite" → acid name ending is 045 ...

9. If the prefix "hydro-" were used to name a ternary acid, what problem would this create when

- naming HCIO? CIO3' has the name of Chlorate. "ate" must be replaced with "ic" If "hydro" is used as a prefix, it would be hydrochloric acid. This has the formula HCI, not HCIO3.

 The prefix "hydro" is not used in the name when an acid contain more than two elements and one of the elements is oxygen.

 "ate is replicate with it and "ite" is replicate with "ous".

Circle the acid(s) below that would be named beginning with the prefix "hydro-."

FESO.

H,CO.

HNO.

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JR Anderson

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