

Nelson Biology VCE Units 3 & 4

Answers to Chapter Questions

Chapter 1

Review questions

- 1 Both non-living and living things are composed of matter. Matter is made up of atoms.
- 2 Living things are capable of reproduction, growth and movement.
- 3 6 electrons
- 4 The proton defines the element.
- 5 The atomic number would be 15 and the mass number would be 32. The symbol for phosphorus is P. The standard notation is $^{32}_{15}\text{P}$.
- 6 The number of neutrons may vary in all the isotopes of a particular element. Hence, an element may have different mass numbers.
- 7 The four major elements found in living things are oxygen, carbon, hydrogen and nitrogen.
- 8 Radioactive isotopes are used as a source of radiation in therapy and as tracers to tag particular compounds.
- 9 A compound is a pure substance made up of more than one type of atom held together by chemical bonds.
- 10 Ionic compounds are composed of cations and anions, which form ionic bonds. Molecular compounds result from atoms sharing their electrons to form covalent bonds. Molecular compounds vary in their solubility in water depending upon their polar nature. Ionic compounds are generally more soluble in water than molecular compounds.
- 11 Molecular compounds result from the sharing of electrons between non-metal elements, whereas ionic compounds are composed of ions and result from the transfer of electrons between metallic and non-metallic elements. As a result of the charges carried, ionic compounds tend to be more readily dissolved in water compared to molecular compounds.
- 12 A chemical bond is the force of attraction holding atoms together in chemical molecules.
- 13 Molecules that have permanent regions of opposite charge are termed 'polar'. Non-polar molecules have no permanent regions of opposite charge.
- 14 Water molecules are highly polar and attract one another, forming hydrogen bonds. Hydrogen bonding determines many of the unique properties of water.
- 15 Hydrophilic molecules, such as alcohol, have regions of opposite charge that readily form hydrogen bonds with water molecules and hence dissolve.

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M Woodhall



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