

Honors Chemistry I
Electromagnetic Spectrum Problems

Constants to know:

$$c = 3.00 \times 10^8 \text{ m/sec}; \quad h = 6.63 \times 10^{-34} \text{ Joule-sec}$$

Units to remember:

$$1 \text{ meter} = 10^9 \text{ Angstrom} = 10^3 \text{ pm} = 10^6 \text{ nm}$$

Formulas: $c = \lambda\nu$; $\Delta E = h\nu$; $\Delta E = hc/\lambda$

Units of frequency are in /sec or sec^{-1} which means there is no unit in the numerator. Frequency is defined as the number of waves/sec.

1. What is the frequency of light if it known to have a wavelength of 455 nm?

$$455 \text{ nm} = 4.55 \times 10^{-7} \text{ m} \qquad 3.00 \times 10^8 = (4.55 \times 10^{-7})(\nu) \quad \nu = \underline{6.59 \times 10^{14} \text{ s}^{-1}}$$

- 2A. What is the wavelength of light (in nm) if it has a frequency of 2.33×10^{14} /sec?

$$3.00 \times 10^8 = (\lambda)(2.33 \times 10^{14}) \qquad \lambda = 1.29 \times 10^{-6} \text{ m} = \underline{1.29 \text{ }\mu\text{m}}$$

- 2B. What is this wavelength in Angstroms (Å)? in picometers (pm)?

$$1.29 \text{ }\mu\text{m} = \underline{12.9 \text{ }\text{\AA}} = \underline{1290 \text{ pm}}$$

- 2C. From the chart given, determine what type of electromagnetic radiation it is?

Xray

3. How much energy is released when an electron drops to the ground state and releases light with a frequency of $5.00 \times 10^{14} \text{ sec}^{-1}$?

$$E = (6.636 \times 10^{-34})(5.00 \times 10^{14}) \qquad \underline{E = 3.32 \times 10^{-19} \text{ J}}$$

4. How much energy (in kJ) is released when one mole of the same type of electron drops back to the ground state?

$$3.32 \times 10^{-19} \text{ J/photon} \times (6.022 \times 10^{23} / 1 \text{ mol}) = 2.00 \times 10^5 \text{ J/mol} = \underline{2.00 \times 10^2 \text{ kJ/mol}}$$

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