Chemistry Unit 8 Worksheet 4 Samples of Every Kind of Problem

On a separate sheet of paper, write a complete solution to each of the problems below. Follow the procedure outlined in class. Be sure to circle your final answer.

 Calculate the number of moles of potassium chlorate, KClO₃ (s), that must decompose to produce potassium chloride, KCl (s), and 1.8 moles of oxygen gas.

Equation:	2 KClO ₃ →		3 O2
Before:	1.2	•	0
Change:	-1.2	+1.2	+1.8
After:	0	1.2	1.8

1.2 moles KClO₂ needed

2. In a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 2.43 g of magnesium?

Equation:	Mg +	2 HCl ->	MgCl ₂ +	H12
Before:	0.100	7	0	O
Change:	-0.100	-0.200 +	0.100 +	0.100
After:	0	-2	0.100	0.100

0.200 moles HCl needed

183 g

3. Ethane, C₂H₆ reacts with oxygen gas to produce carbon dioxide gas and water vapor. What mass of oxygen gas is required to react with 2.20 moles of ethane?

Equation:	2 C2H6	+ 7 O ₂	→ 4 CO;	+ 6 H ₂ O
Before:	2.20	200 200	0	0
Change:	-2.20	- 7.70	+ 4,40	+ 6.60
After:	0	2.5	4.40	6.60

O₂ needed: 7.70 mole x 32.0 g = 246 g O₂ required I mole

 Determine the mass of sodium nitrate produced when 0.73 g of nickel (II) nitrate reacts with sodium hydroxide according to the following equation: Ni(NO₃)₂ + 2 NaOH --> Ni(OH)₂ + 2 NaNO₃

Equation:	Ni(NO ₃) ₂ +	2 NaOH	→ Ni(OH)2 +	2 NaNO3
Before:	0.0040	XS	0	.0
Change:	- 0.0040	- 0.0080	+0.0040	+0.0080
After:	0	XS	0.0040	0.0080

NaNO₃ produced: 0.0080 mole x <u>85.0 g</u> = 0.68 g 1 mole

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