

8. How many moles of copper atoms are needed to supply the same number of atoms as does 235 grams of iron? Key.

$$235 \text{ g} \left(\frac{1 \text{ mol}}{55.8 \text{ g}} \right) = 4.21 \text{ moles of Fe atoms}$$

4.21 mols of Cu atoms would have the same number of atoms.

9. How many grams of carbon are needed to supply twice as many atoms as does 27 grams of beryllium?

$$27 \text{ g} \left(\frac{1 \text{ mol}}{9.0 \text{ g}} \right) = 3.0 \text{ mols Be atoms} \quad \text{twice atoms } 2(3.0 \text{ mols}) = 6.0 \text{ mols}$$

$$6.0 \text{ mols C} \left(\frac{12.0 \text{ g}}{1 \text{ mol}} \right) = 72 \text{ g of carbon.}$$

10. 62.0 grams of rust is composed of 43.4 grams of iron. Determine the empirical formula for rust.

$$43.4 \text{ g Fe} \rightarrow 43.4 \text{ g} \div 55.85 \text{ g/mol} = 0.78 \text{ mol Fe} \quad \text{mol ratio Fe : O}$$

$$62.0 - 43.4 \text{ g O} \rightarrow 18.6 \text{ g} \div 16.0 \text{ g/mol} = 1.16 \text{ mol O} \quad 2 : 3$$



11. A solution of NaCl contains 35 g of solute dissolved in 100g of water. What is the molarity of the solution?

$$\text{FW NaCl} = 58.5 \text{ g/mol} \quad 100 \text{ g H}_2\text{O} \rightarrow 100 \text{ mL H}_2\text{O} \rightarrow 0.100 \text{ L}$$

$$35 \text{ g} \div 58.5 \text{ g/mol} = 0.60 \text{ mol}$$

$$M = \frac{n}{V} = \frac{0.60 \text{ mol}}{0.100 \text{ L}} = 6.0 \text{ M}$$

12. The following three containers have identical capacities. Each contains Avogadro's number of molecules of a gaseous element. Fill in the blanks with the missing information.



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D Keegan



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Table of Contents Nelson Chemistry 11 Answer Key Chapter 8

1. Understanding the eBook Nelson Chemistry 11 Answer Key Chapter 8
 - The Rise of Digital Reading Nelson Chemistry 11 Answer Key Chapter 8
 - Advantages of eBooks Over Traditional Books
2. Identifying Nelson Chemistry 11 Answer Key Chapter 8
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Nelson Chemistry 11 Answer Key Chapter 8
 - User-Friendly Interface
4. Exploring eBook Recommendations from Nelson Chemistry 11 Answer Key Chapter 8

- Personalized Recommendations
 - Nelson Chemistry 11 Answer Key Chapter 8 User Reviews and Ratings
 - Nelson Chemistry 11 Answer Key Chapter 8 and Bestseller Lists
5. Accessing Nelson Chemistry 11 Answer Key Chapter 8 Free and Paid eBooks
- Nelson Chemistry 11 Answer Key Chapter 8 Public Domain eBooks
 - Nelson Chemistry 11 Answer Key Chapter 8 eBook Subscription Services
 - Nelson Chemistry 11 Answer Key Chapter 8 Budget-Friendly Options
6. Navigating Nelson Chemistry 11 Answer Key Chapter 8 eBook Formats
- ePub, PDF, MOBI, and More
 - Nelson Chemistry 11 Answer Key Chapter 8 Compatibility with Devices
 - Nelson Chemistry 11 Answer Key Chapter 8 Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Nelson Chemistry 11 Answer Key Chapter 8
 - Highlighting and Note-Taking Nelson Chemistry 11 Answer Key Chapter 8
 - Interactive Elements Nelson Chemistry 11 Answer Key Chapter 8
8. Staying Engaged with Nelson Chemistry 11 Answer Key Chapter 8
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Nelson Chemistry 11 Answer Key Chapter 8
9. Balancing eBooks and Physical Books Nelson Chemistry 11 Answer Key Chapter 8
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Nelson Chemistry 11 Answer Key Chapter 8
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Nelson Chemistry 11 Answer Key Chapter 8
- Setting Reading Goals Nelson Chemistry 11 Answer Key Chapter 8
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Nelson Chemistry 11 Answer Key Chapter 8

- Fact-Checking eBook Content of Nelson Chemistry 11 Answer Key Chapter 8
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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