

Molecular Chemistry (part of a curriculum)

Model 1: write symbols for three isotopes of carbon



Critical Thinking Questions

- How many protons are in carbon-12? 6 How many neutrons are in carbon-12? 6
- How many protons are in carbon-13? 6 How many neutrons are in carbon-13? 7
- How many protons are in carbon-14? 6 How many neutrons are in carbon-14? 8
- What is similar about the three and what different about them?

Students must think critically, consider, assess, evaluate, and synthesize their knowledge of atomic structure, isotopes, and the periodic table, and apply it to the problem.

- What does the subscript in isotopes tell about the number of protons? 6 What does the superscript tell about the number of neutrons? 12

Students must be able to identify all protons, neutrons, and electrons in an atom, and be able to calculate the mass of an atom.

A student response is a correct answer to the question of the mass of an atom. This is not a "correct" answer to the question of the mass of an atom, but it is a correct answer to the question of the mass of an atom.

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Table 1: Isotopes of carbon and isotopes of carbon-12 produced by nuclear reactions

Isotope of carbon	Symbol	Mass number	Atomic number	Relative atomic mass	Relative abundance	Relative mass
Carbon-12	$^{12}_6\text{C}$	12	6	12.000	98.9%	12.000
Carbon-13	$^{13}_6\text{C}$	13	6	13.003	1.1%	13.003
Carbon-14	$^{14}_6\text{C}$	14	6	14.003	0.0001%	14.003
Carbon-15	$^{15}_6\text{C}$	15	6	15.003	0.0001%	15.003
Carbon-16	$^{16}_6\text{C}$	16	6	16.003	0.0001%	16.003

Nuclear Chemistry Chemactivity 8 Answers

Didier Musso



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