

Nuclear Chemistry Worksheet 3

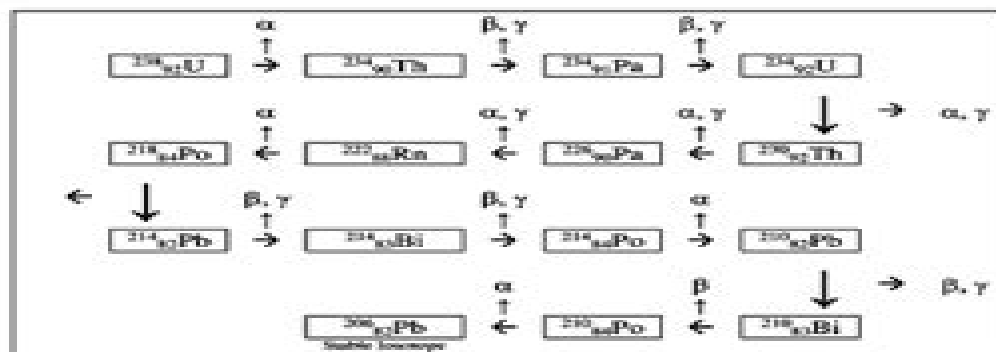
Directions: Identify the following as alpha, beta, gamma, or neutron.

- $\frac{1}{0}n$ _____
- $\frac{0}{-1}e$ _____
- $\frac{4}{2}\text{He}$ _____
- $\frac{0}{0}\gamma$ _____
- Nuclear decay with no mass and no charge _____
- An electron _____
- Least penetrating nuclear decay _____
- Most damaging nuclear decay to the human body _____
- Nuclear decay that can be stopped by skin or paper. _____
- Nuclear decay that can be stopped by aluminum. _____

Complete the following nuclear equations.

- ${}_{19}^{42}\text{K} \rightarrow {}_{-1}^0e + \underline{\hspace{2cm}}$
- ${}_{94}^{239}\text{Pu} \rightarrow \frac{4}{2}\text{He} + \underline{\hspace{2cm}}$
- ${}_{4}^9\text{Be} \rightarrow \frac{9}{4}\text{Be} + \underline{\hspace{2cm}}$
- ${}_{92}^{235}\text{U} \rightarrow \underline{\hspace{2cm}} + {}_{90}^{231}\text{Th}$
- ${}_{3}^6\text{Li} \rightarrow \frac{4}{2}\text{He} + \underline{\hspace{2cm}}$
- $\underline{\hspace{2cm}} \rightarrow {}_{56}^{142}\text{Ba} + {}_{36}^{91}\text{Kr} + 3\frac{1}{0}n$

Nuclear Decay Series



The figure maps the radioactive decay of uranium-238 to lead-206. Use the figure to answer the following questions.

- How many alpha particles are produced as one atom of uranium-238 decays to an atom of lead-206? _____
- How many beta particles? _____
- Explain why lead-206 is a stable isotope. _____

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