

Nuclear chemistry help sheet

1. Atomic Number and Mass Number

- a. Atomic Notation – commonly used to specify the number of protons and neutrons in the atom

Mass number 235 Atomic symbol
Atomic number $_{92}$ U

- b. Atomic Number – the number of protons
c. Mass Number – the number of protons + neutrons
d. Number of neutrons = mass number – atomic number
e. Isotopes – atoms of an element that have different number of neutrons.
 ${}^1\text{H}$, ${}^2\text{H}$, and ${}^3\text{H}$ ${}^{31}\text{P}$ and ${}^{32}\text{P}$

2. Common forms of radioactive decay

Decay Type	Particle	Particle Mass	Particle Charge	Change in mass number	Change in Atomic number
Alpha	${}^4_2\text{He}$ or ${}^4_2\alpha$	4	2+	Decreases by 4	Decreases by 2
Beta	${}^0_{-1}\beta$ or ${}^0_{-1}e^-$	0	1-	No change	Increases by 1
Positron	${}^0_{+1}e$ or ${}^0_{+1}\beta$	0	1+	No change	Decreases by 1
Electron Capture	${}^0_{-1}\beta$ or ${}^0_{-1}e^-$	0	1-	No change	Decreases by 1
Gamma	${}^0_0\gamma$	0	0	No Change	No Change

3. Balancing nuclear equation

- Sum of reactants mass = sum of products mass
- Sum of reactants atomic numbers = sum of product atomic numbers
- Elements may change
- Example:



4. Half-lives and isotopic dating

- a. Radioactive decay follows first order kinetics (see kinetics handout).
- b. Half-life is characteristic of each isotope
- c. Relative abundance is used estimate the age of objects
 - i. Age of rock – U-235 – half-life of 4.5 billion years
 - ii. Organic materials in archeological sites – C-14 – half-life of 5730 years - Dates between 500 and 50,000 years old
 - iii. Age of wines – H(T)-3 – half-life of 12.26 years – dates up to 100 years

5. Fusion, Fission, Transmutation

- Fusion – smaller atoms come together to form larger atoms
- Fission – larger atoms break apart to form small atoms
- Transmutation (artificial) – elements are converted from one to another

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JS Bruner



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