

Focus 3: States of Matter

3A: The Nature of Gases

11. Pressure is the amount of force exerted by a gas over a certain area.
- $P = F/A$
 - The SI unit of pressure is the pascal, where $1 \text{ Pa} = 1 \text{ kg/m}^2 \cdot \text{s}^2$
 - $1 \text{ bar} = 1.013 \times 10^5 \text{ Pa}$
 - $1 \text{ atm} = 1.01325 \times 10^5 \text{ Pa}$
 - $1 \text{ atm} = 760 \text{ Torr}$

3B: The Gas Laws

8. Boyle's Law states that, for a fixed amount of gas at a constant temperature, the volume of the gas is inversely proportional to the pressure.
9. Boyle's Law can be expressed as $P_1V_1 = P_2V_2$.
10. Avogadro's Principle states that, at constant temperature and pressure, the volume of a gas is directly proportional to the number of moles of gas.
11. The Ideal Gas Law is expressed as $PV = nRT$.
12. The units for the gas constant R are $\text{L}\cdot\text{atm}/\text{mol}\cdot\text{K}$.
13. The units for the gas constant R are $\text{J}/\text{mol}\cdot\text{K}$.
14. The units for the gas constant R are $\text{cal}/\text{mol}\cdot\text{K}$.
15. The units for the gas constant R are $\text{m}^3\cdot\text{Pa}/\text{mol}\cdot\text{K}$.
16. The units for the gas constant R are $\text{L}\cdot\text{bar}/\text{mol}\cdot\text{K}$.
17. The units for the gas constant R are $\text{L}\cdot\text{mmHg}/\text{mol}\cdot\text{K}$.
18. The units for the gas constant R are $\text{L}\cdot\text{torr}/\text{mol}\cdot\text{K}$.
19. The units for the gas constant R are $\text{L}\cdot\text{atm}/\text{mol}\cdot\text{K}$.
20. The units for the gas constant R are $\text{L}\cdot\text{atm}/\text{mol}\cdot\text{K}$.

TABLE 28.1 The Gas Constant, R

$\mu_{\text{H}^+} = 7.44 \times 10^{-2} \text{ eV/mass } \text{H}^+ \text{ (measured)}$
 $\mu_{\text{H}^+} = 4.61 \times 10^{-2} \text{ eV/mass } \text{H}^+ \text{ (measured)}$
 $\mu_{\text{H}^+} = 4.65 \text{ eV/mass } \text{H}^+ \text{ (measured)}$
 $\mu_{\text{H}^+} = 4.65 \text{ eV/mass } \text{H}^+ \text{ (measured)}$
 $\mu_{\text{H}^+} = 2.26 \text{ eV/mass } \text{H}^+ \text{ (measured)}$

$$\frac{1}{2} \left(\frac{1}{2} + \frac{1}{2} \right) = \frac{1}{2}$$

- STP = standard temperature and pressure (0 degrees C and 1 atm)

Sample Problems 14b Answer Key

DJ Losen



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