

**Directions:** Choose the letter that corresponds to the correct answer.

1. Which of the following can be considered as Dobereiner's triad?
  - a. He: 4.0, Li: 6.9, Na: 23.0
  - b. N: 14.0, P: 40.0, At: 217.0
  - c. Ne: 20.2, Ar: 39.9, Kr: 83.8
  - d. Ca: 40.1, Sr: 87.6, Ba: 137.3
2. Which of the following elements does not belong to the group?
  - a. N
  - b. S
  - c. Po
  - d. Te
3. Which of the following is/are TRUE about boron and oxygen?
  - a. both are metalloids
  - b. boron is smaller than oxygen
  - c. boron is more electronegative than oxygen
  - d. higher energy is required to eject an electron from the valence shell of boron
  - e. all of the above
  - f. none of the above

For 4 and 5, consider Na, Mg, Al, and Si.

4. The element that ejects one electron from its valence shell most readily is \_\_\_\_\_.
  - a. Na
  - b. Mg
  - c. Al
  - d. Si
5. Which element has the most negative electron affinity?
  - e. Na
  - f. Mg

# Periodic Trends Practice Questions

**The Princeton Review**



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