

Percent Yield Practice

3 pages + 3-page Answer KEY + Periodic Table

Chemistry

PERCENT YIELD

Show all work and circle or highlight your final answer.

$$\text{percent yield (\%)} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$$

- 1 16.8 grams of sodium reacts with excess chlorine gas.



The theoretical yield of sodium chloride is 25.5 grams. The actual yield was 23.5 grams. Calculate the percent yield of sodium chloride.

- 2 4.8 grams of carbon reacts with excess oxygen gas.



The theoretical yield of carbon dioxide is 25.5 g. Calculate the percent yield of carbon dioxide.

- 3 1.6 grams of zinc reacts with excess sulfur.



The theoretical yield of zinc sulfide is 15.5 grams. Calculate the percent yield of zinc sulfide.

- 4 1.0 grams of phosphorus reacts with excess chlorine.



The theoretical yield of phosphorus trichloride is 4.5 grams. Calculate the percent yield of phosphorus trichloride.

- 5 Hydrogen gas (H_2) reacts with oxygen gas (O_2) to produce water (H_2O). 5.8 grams of hydrogen reacts with excess oxygen gas.



What is the theoretical yield?

- 6 The actual yield was 16.8 grams. Calculate the percent yield.

- 7 Magnesium (Mg) reacts with hydrochloric acid (HCl) to produce magnesium chloride (MgCl_2) and hydrogen gas (H_2). 1.0 gram of magnesium reacts with excess hydrochloric acid.



What is the theoretical yield?

- 8 The actual yield was 40.2 grams. Calculate the percent yield.

- 9 Calcium carbonate (CaCO_3) reacts with hydrochloric acid (HCl) to produce calcium chloride (CaCl_2), carbon dioxide (CO_2), and water (H_2O). 15.8 grams of calcium carbonate reacts with excess hydrochloric acid.



What is the theoretical yield of calcium chloride?

$$\begin{array}{r} 15.8 \text{ g CaCO}_3 \times \frac{1 \text{ mol CaCO}_3}{100.09 \text{ g CaCO}_3} \times \frac{1 \text{ mol CaCl}_2}{1 \text{ mol CaCO}_3} \times \frac{146.97 \text{ g CaCl}_2}{1 \text{ mol CaCl}_2} \\ = 0.158 \text{ mol CaCO}_3 \times 1 \text{ mol CaCl}_2 \\ = 0.158 \text{ mol CaCl}_2 \end{array}$$

$$0.158 \text{ mol CaCl}_2 \times \frac{1 \text{ mol CaCl}_2}{1 \text{ mol CaCl}_2} \times \frac{146.97 \text{ g CaCl}_2}{1 \text{ mol CaCl}_2} = 23.12 \text{ g CaCl}_2$$

$$\begin{array}{r} = 0.158 \text{ mol CaCl}_2 \times \frac{146.97 \text{ g CaCl}_2}{1 \text{ mol CaCl}_2} \\ = 23.12 \text{ g CaCl}_2 \end{array}$$

Molar Mass CaCl_2

$$\text{Ca} = 40.08$$

$$\text{Cl} = 35.45$$

$$\text{Cl} = 35.45 \times 2 = 70.90$$

$$\text{Molar Mass} = 40.08 + 70.90 = 110.98 \text{ g}$$

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$$\text{Molar Mass} = 40.08 + 70.90 = 110.98 \text{ g}$$

$$\text{Molar Mass} = 110.98 \text{ g}$$

- 10 The actual yield was 23.5 grams. Calculate the percent yield of calcium chloride.

$$\text{percent yield (\%)} = \frac{23.5 \text{ g}}{23.12 \text{ g}} \times 100 = 1.0164 \times 100 = 101.64\% \text{ CaCl}_2$$

- 11 Potassium (K) reacts with iodine (I_2) to produce potassium iodide (KI). 1.8 grams of potassium reacts with excess iodine.



What is the theoretical yield?

$$\begin{array}{r} 1.8 \text{ g K} \times \frac{1 \text{ mol K}}{39.10 \text{ g K}} \times \frac{2 \text{ mol KI}}{2 \text{ mol K}} \times \frac{166.10 \text{ g KI}}{1 \text{ mol KI}} \\ = 0.046 \text{ mol K} \times 2 \text{ mol KI} \\ = 0.092 \text{ mol KI} \end{array}$$

Molar Mass KI

$$\text{K} = 39.10$$

$$\text{I} = 126.90$$

$$\text{Molar Mass} = 39.10 + 126.90 = 166.00 \text{ g}$$

$$\text{Molar Mass} = 166.00 \text{ g}$$

$$\text{Molar Mass} = 166.00 \text{ g}$$

$$\text{Molar Mass} = 166.00 \text{ g}$$

$$\text{Molar Mass} = 166.00 \text{ g}$$

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$$\text{Molar Mass} = 166.00 \text{ g}$$

Percent Yield Example Answer Key

CH Cherryholmes

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'CH Cherryholmes' text.

Percent Yield Example Answer Key:

Percent Yield Example Answer Key: Bestsellers in 2023 The year 2023 has witnessed a noteworthy surge in literary brilliance, with numerous compelling novels captivating the hearts of readers worldwide. Lets delve into the realm of top-selling books, exploring the fascinating narratives that have charmed audiences this year. Percent Yield Example Answer Key : Colleen Hoover's "It Ends with Us" This poignant tale of love, loss, and resilience has captivated readers with its raw and emotional exploration of domestic abuse. Hoover skillfully weaves a story of hope and healing, reminding us that even in the darkest of times, the human spirit can triumph. Uncover the Best : Taylor Jenkins Reids "The Seven Husbands of Evelyn Hugo" This spellbinding historical fiction novel unravels the life of Evelyn Hugo, a Hollywood icon who defies expectations and societal norms to pursue her dreams. Reids captivating storytelling and compelling characters transport readers to a bygone era, immersing them in a world of glamour, ambition, and self-discovery. Percent Yield Example Answer Key : Delia Owens "Where the Crawdads Sing" This mesmerizing coming-of-age story follows Kya Clark, a young woman who grows up alone in the marshes of North Carolina. Owens crafts a tale of resilience, survival, and the transformative power of nature, captivating readers with its evocative prose and mesmerizing setting. These bestselling novels represent just a fraction of the literary treasures that have emerged in 2023. Whether you seek tales of romance, adventure, or personal growth, the world of literature offers an abundance of engaging stories waiting to be discovered. The novel begins with Richard Papen, a bright but troubled young man, arriving at Hampden College. Richard is immediately drawn to the group of students who call themselves the Classics Club. The club is led by Henry Winter, a brilliant and charismatic young man. Henry is obsessed with Greek mythology and philosophy, and he quickly draws Richard into his world. The other members of the Classics Club are equally as fascinating. Bunny Corcoran is a wealthy and spoiled young man who is always looking for a good time. Charles Tavis is a quiet and reserved young man who is deeply in love with Henry. Camilla Macaulay is a beautiful and intelligent young woman who is drawn to the power and danger of the Classics Club. The students are all deeply in love with Morrow, and they are willing to do anything to please him. Morrow is a complex and mysterious figure, and he seems to be manipulating the students for his own purposes. As the students become more involved with Morrow, they begin to commit increasingly dangerous acts. The Secret History is a exceptional and suspenseful novel that will keep you guessing until the very end. The novel is a cautionary tale about the dangers of obsession and the power of evil.

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