# Finding the Percent Composition

Percent Composition - The percentage (by mass) of each element in a compound.

#### Example 1:

Find the percent composition of red candies in a bag of 200 candies, if there are 26 red, and 174 blue candies.

Total Candies: 200

Red Candies: 26

Blue Candies: 174

26 red	× 100	
200 total	A IUU	

Formula:

Part x 100 Total

#### Example 2:

Find the percent composition of MgCl<sub>2</sub>.

- Total mass MgCl2: 95.21 g
- Mass of Mg: 24.31 g
- 70.91 g Mass of Cl2:

#### Example 3:

Find the percent composition of H2SO4.

- Total mass H<sub>2</sub>SO<sub>4</sub>: 98.08 g
- 2.02 g Mass of H<sub>2</sub>:
- Mass of S: 32.07 g
- Mass of O4: 64.00 g

$$\frac{2.02 \text{ g}}{98.08 \text{ g}} \times 100$$

$$\frac{2.02 \,\mathrm{g}}{98.08 \,\mathrm{g}} \times 100 = 2.06 \,\% \,\mathrm{H}_2$$

# **Percent Composition Answers**

**L** Cohen

**Percent Composition Answers:** 

The book delves into Percent Composition Answers. Percent Composition Answers is an essential topic that must be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Percent Composition Answers, encompassing both the fundamentals and more intricate discussions.

- 1. This book is structured into several chapters, namely:
  - Chapter 1: Introduction to Percent Composition Answers
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  - Chapter 3: Percent Composition Answers in Everyday Life
  - Chapter 4: Percent Composition Answers in Specific Contexts
  - ∘ Chapter 5: Conclusion
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