

Electron Configuration Practice - Homework - **KEY**

In the space below, write the expanded electron configurations of the following elements:

- | | | |
|----|-----------|--|
| 1) | sodium | $1s^2 2s^2 2p^6 3s^1$ |
| 2) | potassium | $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$ |
| 3) | chlorine | $1s^2 2s^2 2p^6 3s^2 3p^5$ |
| 4) | bromine | $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$ |
| 5) | oxygen | $1s^2 2s^2 2p^4$ |

In the space below, write the abbreviated electron configurations of the following elements:

- | | | |
|-----|-----------|-------------------------|
| 6) | manganese | $[\text{Ne}] 3s^2$ |
| 7) | silver | $[\text{Kr}] 5s^2 4d^9$ |
| 8) | nitrogen | $[\text{He}] 2s^2 2p^3$ |
| 9) | sulfur | $[\text{Ne}] 3s^2 3p^4$ |
| 10) | argon | $[\text{Ne}] 3s^2 3p^6$ |

In the space below, write the orbital notation of the following elements:

- | | | | | | | | | | | | | | | | | | |
|---------------|-------------------|----------------------------|----------------------------|----------------------------|----------------------------|----------------------------|----------------------------|----------------------------|----------------------------|----------------------------|----------------------------|----------------------------|----------------------------|----------------------------|----------------------------|----------------------------|------------------|
| 11) manganese | 12 e ⁻ | $\uparrow\downarrow$
1s | $\uparrow\downarrow$
2s | $\uparrow\downarrow$
2p | $\uparrow\downarrow$
2p | $\uparrow\downarrow$
2p | $\uparrow\downarrow$
3s | | | | | | | | | | |
| 12) gallium | 31 e ⁻ | $\uparrow\downarrow$
1s | $\uparrow\downarrow$
2s | $\uparrow\downarrow$
2p | $\uparrow\downarrow$
2p | $\uparrow\downarrow$
2p | $\uparrow\downarrow$
3s | $\uparrow\downarrow$
3p | $\uparrow\downarrow$
3p | $\uparrow\downarrow$
3p | $\uparrow\downarrow$
4s | $\uparrow\downarrow$
3d | $\uparrow\downarrow$
3d | $\uparrow\downarrow$
3d | $\uparrow\downarrow$
3d | $\uparrow\downarrow$
3d | \uparrow
4p |
| 13) nitrogen | 7 e ⁻ | $\uparrow\downarrow$
1s | $\uparrow\downarrow$
2s | \uparrow
2p | \uparrow
2p | \uparrow
2p | | | | | | | | | | | |
| 14) sulfur | 16 e ⁻ | $\uparrow\downarrow$
1s | $\uparrow\downarrow$
2s | $\uparrow\downarrow$
2p | $\uparrow\downarrow$
2p | $\uparrow\downarrow$
2p | $\uparrow\downarrow$
3s | $\uparrow\downarrow$
3p | \uparrow
3p | \uparrow
3p | | | | | | | |
| 15) argon | 18 e ⁻ | $\uparrow\downarrow$
1s | $\uparrow\downarrow$
2s | $\uparrow\downarrow$
2p | $\uparrow\downarrow$
2p | $\uparrow\downarrow$
2p | $\uparrow\downarrow$
3s | $\uparrow\downarrow$
3p | $\uparrow\downarrow$
3p | $\uparrow\downarrow$
3p | | | | | | | |

Determine what elements are denoted by the following electron configurations:

- | | | |
|-----|---|-------------|
| 16) | $1s^2 2s^2 2p^6 3s^2 3p^4$ | sulfur |
| 17) | $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$ | rubidium |
| 18) | $[\text{Kr}] 5s^2 4d^{10} 5p^3$ | antimony |
| 19) | $[\text{Xe}] 6s^2 4f^{14} 5d^6$ | osmium |
| 20) | $[\text{Rn}] 7s^2 5f^{11}$ | einsteinium |

Determine which of the following electron configurations are not valid:

- | | | |
|-----|--|--|
| 21) | $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^5$ | not valid (take a look at "4d") |
| 22) | $1s^2 2s^2 2p^6 3s^3 3d^5$ | not valid (3p comes after 3s) |
| 23) | $[\text{Ra}] 7s^2 5f^6$ | not valid (radium isn't a noble gas) |
| 24) | $[\text{Kr}] 5s^2 4d^{10} 5p^5$ | valid |
| 25) | $[\text{Xe}]$ | not valid (an element can't be its own electron configuration) |

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1. Understanding the eBook Science Electron Configuration Packet Answer Key
 - The Rise of Digital Reading Science Electron Configuration Packet Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Science Electron Configuration Packet Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Science Electron Configuration Packet Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Science Electron Configuration Packet Answer Key
 - Personalized Recommendations
 - Science Electron Configuration Packet Answer Key User Reviews and Ratings
 - Science Electron Configuration Packet Answer Key and Bestseller Lists
5. Accessing Science Electron Configuration Packet Answer Key Free and Paid eBooks
 - Science Electron Configuration Packet Answer Key Public Domain eBooks
 - Science Electron Configuration Packet Answer Key eBook Subscription Services
 - Science Electron Configuration Packet Answer Key Budget-Friendly Options

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 - ePub, PDF, MOBI, and More
 - Science Electron Configuration Packet Answer Key Compatibility with Devices
 - Science Electron Configuration Packet Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Science Electron Configuration Packet Answer Key
 - Highlighting and Note-Taking Science Electron Configuration Packet Answer Key
 - Interactive Elements Science Electron Configuration Packet Answer Key
8. Staying Engaged with Science Electron Configuration Packet Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Science Electron Configuration Packet Answer Key
9. Balancing eBooks and Physical Books Science Electron Configuration Packet Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Science Electron Configuration Packet Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Science Electron Configuration Packet Answer Key
 - Setting Reading Goals Science Electron Configuration Packet Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Science Electron Configuration Packet Answer Key
 - Fact-Checking eBook Content of Science Electron Configuration Packet Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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