

Unit 1: Chemistry

Notes

TEST

E/Us	/20	T/5	/10	COMM	/15	APP	/5
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Part 1: Multiple Choice

Mark the appropriate letter on the grid below (20 marks- E/Us).

1. Which of the following is an example of a physical change?

- a) Energy is released or absorbed
- b) A precipitate forms
- c) A gas is produced
- d) Light is emitted
- e) Change of state

2. Which of the following elements is an example of a halogen?

- a) Oxygen
- b) Neon
- c) Bromine
- d) Lithium
- e) Magnesium

3. Which of the following states of matter indicates that a substance is dissolved in water?

- a) (s)
- b) (l)
- c) (g)
- d) (aq)
- e) (mg)

4. In all chemical reactions the:

- a) the mass and volume of both the reactants and the products must be equal.
- b) the masses of the combining chemicals must be equal.
- c) total mass of the reactants equals the total mass of the products
- d) the total volume of the reactants equals the total volume of the products

5. Al₂O₃ _____ is a pure substance that cannot be broken down further by physical or

chemical means.

- a) H₂O
- b) H₂SO₄
- c) CaO
- d) H₂O
- e) CO₂

6. When a chemical reaction takes place in a closed system, which of the following statements best describes the principle of conservation of mass?

- a) The total mass of the substances before the reaction is greater than the total mass of the substances after the reaction.
- b) The total mass of the substances before the reaction is equal to the total mass of the substances after the reaction.
- c) The total mass of the substances before the reaction is less than the total mass of the substances after the reaction.
- d) The total mass of the reactants decreases, while the total mass of the products increases.

7. What type of reaction occurs between calcium and hydrochloric acid (HCl)?

- a) synthesis
- b) decomposition
- c) single displacement
- d) double displacement

8. _____ is the deterioration of any metal.

- a) Oxidation

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AW Rasmussen



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