

Molecular Polarity

How do we know if one molecule will be attracted to another molecule?

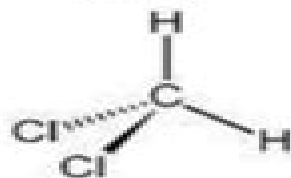
Why?

Did you ever wonder how molecules of a substance stay close together? Why does water form droplets or support things on its surface? Electrons may be distributed evenly or unevenly throughout a molecule, which creates partial charges at different parts of the molecule. These partial charges on one molecule often interact with the partial charges of a neighboring molecule. In this activity you will learn how to determine molecular polarity and placement of partial charges molecules.

Model 1 – The Molecules

POLAR MOLECULES:

Tetrahedral



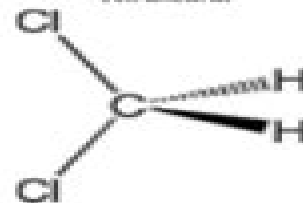
(CH₂Cl₂)

Trigonal Pyramidal



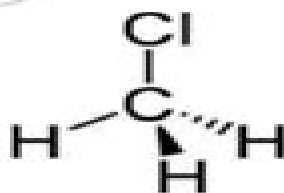
(NH₃)

Tetrahedral



(CH₃Cl)

Tetrahedral



(CH₃Cl)

Bent



(H₂O)

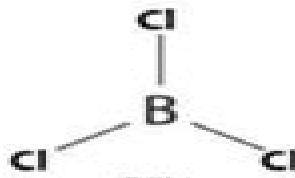
NON-POLAR MOLECULES:

Tetrahedral



(CCl₄)

Trigonal Planar



(BCl₃)

Tetrahedral



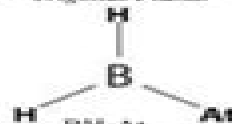
(CH₄)

Linear



(CO₂)

Trigonal Planar



(BH₃)

Bond Type Guide:

- C – Cl = Polar Covalent
- C – H = Polar Covalent
- C – O = Polar Covalent
- N – H = Polar Covalent
- O – H = Polar Covalent
- B – Cl = Polar Covalent
- B – H = Non-Polar Covalent
- B – Al = Non-Polar Covalent

Polarity Of Molecules Pogil

LP Steffe

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'LP Steffe' text.

Polarity Of Molecules Pogil:

Polarity Of Molecules Pogil Book Review: Unveiling the Power of Words

In a global driven by information and connectivity, the power of words has be more evident than ever. They have the ability to inspire, provoke, and ignite change. Such is the essence of the book **Polarity Of Molecules Pogil**, a literary masterpiece that delves deep in to the significance of words and their affect our lives. Compiled by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we will explore the book is key themes, examine its writing style, and analyze its overall effect on readers.

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