

Colligative properties are properties that depend only upon the number of solute atoms, ions, or molecules in a solution and not on the nature of those atoms, ions or molecules. Freezing point depression and boiling point elevation are examples of colligative properties. Raoult discovered that the addition of solute particles causes the boiling point of a solution to be elevated and the freezing point to be depressed. It has been found that 1 mole of nonvolatile, nonionizing solute particles will raise the boiling point of 1.0 kg of water by 0.52 °C, to 100.52°C at 1 atm of pressure. The same concentration will lower the freezing point of 1.0 kg of water by 1.86°C, to -1.86°C. These two numbers apply only to water and are called the molal boiling point and freezing point constants for water. Other solvents have different constants.

PROPERTY	BOILING POINT ELEVATION	FREEZING POINT DEPRESSION
EQUATION	Boiling Point Elevation: A solution will boil at a higher temperature than the pure solvent.	Freezing Point Depression: A solution will freeze at a lower temperature than the pure solvent.
	$\Delta T_b = i K_b m$	$\Delta T_f = i K_f m$
	ΔT_b = change in boiling temp. in (°C) <i>i</i> = a unitless constant (called the van't Hoff factor) associated with the degree of dissociation of the solute in a solvent. K_b = molal boiling point elevation constant (°C/m) For H₂O: K_b = .52°C/m <i>m</i> = molality of the solution	ΔT_f = change in freezing temp. in (°C) <i>i</i> = a unitless constant (called the van't Hoff factor) associated with the degree of dissociation of the solute in a solvent. K_f = molal freezing point lowering constant (°C/m) For H₂O: K_f = 1.86°C/m <i>m</i> = molality of the solution
GUIDELINES	<i>i</i> = 1 for substances that do not ionize 2 for substances which ionize into two ions like LiCl 3 for substances which ionize into three ions like CaCl ₂ ...etc. <i>i</i> may have to be calculated if a % of dissociation is given for a solute. If the solvent is not water you must find its K_b .	<i>i</i> = 1 for substances that do not ionize 2 for substances which ionize into two ions like LiCl 3 for substances which ionize into three ions like CaCl ₂ ...etc. <i>i</i> may have to be calculated if a % of dissociation is given for a solute. If the solvent is not water you must find its K_f .

EXAMPLE #1

If 90.0 g of nonionizing C₆H₁₂O₆, are dissolved in 255 g of H₂O, what will be the boiling point of the resulting solution?

First calculate the molality (*m*) of the solution:

$$m = \frac{92\text{g}}{\left(180 \frac{\text{g}}{\text{mole}}\right)(0.255 \text{ kg of solvent})} = 2.0 \text{ m}$$

Next apply the boiling point elevation equation:

$$[\textit{i} = 1 \text{ for a nonionizing solute.}]$$

$$\Delta T_b = i K_b m$$

$$\Delta T_b = (1)(0.52^\circ\text{C/m})(2.0 \text{ m})$$

$$\Delta T_b = 1.04^\circ\text{C}$$

$$\text{BP} = (\Delta T_b + 100.00^\circ\text{C})$$

$$\text{BP} = (1.04^\circ\text{C} + 100.00^\circ\text{C}) = \underline{101.04^\circ\text{C}}$$

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N Noddings



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Modern Chemistry Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001

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Chemistry Morris Hein, Susan Arena, Cary Willard, 2023 Foundations of College Chemistry 16th edition presents chemistry as a modern vital subject and is designed to make introductory chemistry accessible to all beginning students. It is intended for students who have never taken a chemistry course or those who had a significant interruption in their studies but plan to continue with the general chemistry sequence. The central focus is to make chemistry interesting and understandable and teach students the problem solving skills they will need. This International Adaptation offers new and updated content with improved presentation of all course material. It builds on the strengths of previous editions including clear explanations and step by step problem solving. The material emphasizes real world applications of chemistry as the authors develop the principles that form the foundation for the further study of chemistry. There is new and expanded coverage of polarizing power and polarizability, Fajans rules, collision number and mean free path, abnormal molecular masses and van t Hoff factor and applications of radioactivity.

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Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text. This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises. New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application.

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Review Alexander Stone Macnow, 2016-07-05 The most efficient learning for the MCAT results you want. Kaplan's MCAT General Chemistry Review has all the information and strategies you need to score higher on the MCAT. This book features more practice than any other guide plus targeted subject review questions, opportunities for self analysis, a complete online center and thorough instruction on all of the general chemistry concepts necessary for MCAT success from the creators of the 1 MCAT prep course. page 4 of cover

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2018-2019 Kaplan Test Prep, 2017-07-04 Kaplan's MCAT Complete 7 Book Set Subject Review has all the information and strategies you need to score higher on the MCAT. These books feature more practice than any other guide plus targeted strategy review opportunities for self analysis and thorough information on all of the critical thinking skills necessary for

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Instructor's Manual for General Chemistry, Experiment and Theory Richard H. Eastman, 1970

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