

balance nuclear reaction equations

aluminum-26 undergoes positron emission
potassium-40 undergoes beta decay (electron)
uranium-238 undergoes alpha decay



electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



NOMENCLATURE [naming compounds, writing formulas]

Write the formula for each of the following compounds:

- a. Sulfur hexafluoride SF_6
- b. Lithium nitride Li_3N
- c. Chromium (III) carbonate $\text{Cr}_2(\text{CO}_3)_3$
- d. Tin (II) chloride SnCl_2
- e. Ammonium acetate $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$)
- f. Mercury (I) chloride Hg_2Cl_2
- g. Potassium bromate KBrO_3
- h. Hydrosulfuric acid $\text{H}_2\text{S(aq)}$
- i. Chloric acid HClO_3
- j. Sulfurous acid H_2SO_3

Name each of the following compounds.

- a. CuSO_4 Copper (II) sulfate
- b. AlF_3 aluminum fluoride
- c. HI hydrogen iodide
- d. NO nitrogen monoxide
- e. $\text{H}_2\text{Se(aq)}$ hydroselenic acid
- f. HNO_3 nitric acid
- g. HNO_2 nitrous acid
- h. NaHSO_4 sodium hydrogen sulfate

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?

$$40 \xrightarrow{1} 20 \xrightarrow{2} 10 \xrightarrow{3} 5 \xrightarrow{4} 2.5 \xrightarrow{5} 1.25$$

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

$$3 \text{ half lives} \quad 100 \xrightarrow{1} 50 \xrightarrow{2} 25 \xrightarrow{3} 12.5 \text{ mg}$$

Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{array}{rcl} 63 & \times & .692 = 43.596 \rightarrow 44 \\ 65 & \times & .308 = 20.02 \rightarrow 20. \\ & + & \\ & \hline & & 63.62 \rightarrow 64 \end{array}$$

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text does not feature Physiology For more information about Biology for the Informed Citizen with Physiology please search for ISBN 9780195381993 *Foundations of College Chemistry* Morris Hein, Susan Arena, Cary Willard, 2016-08-02 This text is an unbound three hole punched version Used by over 750 000 students Foundations of College Chemistry Binder Ready Version 15th Edition is praised for its accuracy clear no nonsense approach and direct writing style Foundations direct and straightforward explanations focus on problem solving making it the most dependable text on the market Its comprehensive scope proven track record outstanding in text examples and problem sets were all designed to provide instructors with a solid text while not overwhelming students in a difficult course Foundations fits into the prep intro chemistry courses which often include a wide mix of students from science majors not yet ready for general chemistry allied health students in their 1st semester of a GOB sequence science education students for elementary school teachers to the occasional liberal arts student fulfilling a science requirement Foundations was specifically designed to meet this wide array of needs *Green Organic Chemistry and its Interdisciplinary Applications* Vera M. Kolb, 2017-04-21 Green Organic Chemistry and Its Interdisciplinary Applications covers key developments in green chemistry and demonstrates to students that the developments were most often the result of innovative thinking Using a set of selected experiments all of which have been performed in the laboratory with undergraduate students it demonstrates how to optimize and develop green experiments The book dedicates each chapter to individual applications such as Engineering The chemical industry The pharmaceutical industry Analytical chemistry Environmental chemistry Each chapter also poses questions at the end with the answers included By focusing on both the interdisciplinary applications of green chemistry and the innovative thinking that has produced new developments in the field this book manages to present two key messages in a manner where they reinforce each other It provides a single and concise reference for chemists instructors and students for learning about green organic chemistry and its great and ever expanding number of applications McGraw-Hill's SAT Subject Test Chemistry, 3rd Edition Thomas Evangelist, 2011-02-03 Expert guidance on the Chemistry exam Many colleges and universities require you to take one or more SAT II Subject Tests to demonstrate your mastery of specific high school subjects McGraw Hill's SAT Subject Test Chemistry is written by experts in the field and gives you the guidance you need perform at your best This book includes 4 full length sample tests updated for the latest test format 40 top tips to remember on test day Glossary of tested chemistry terms and formulas Tips and strategies from one of the most popular teachers at the renowned Brooklyn Technical High School Diagnostic test to pinpoint strengths and weaknesses Step by step review of all topics covered on the exam In depth coverage of the lab experiment questions that are a major test feature Charts tables and illustrations to simplify and reinforce learning Practice tests just like the real SAT Subject Test in Chemistry Test taking tips and strategies

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