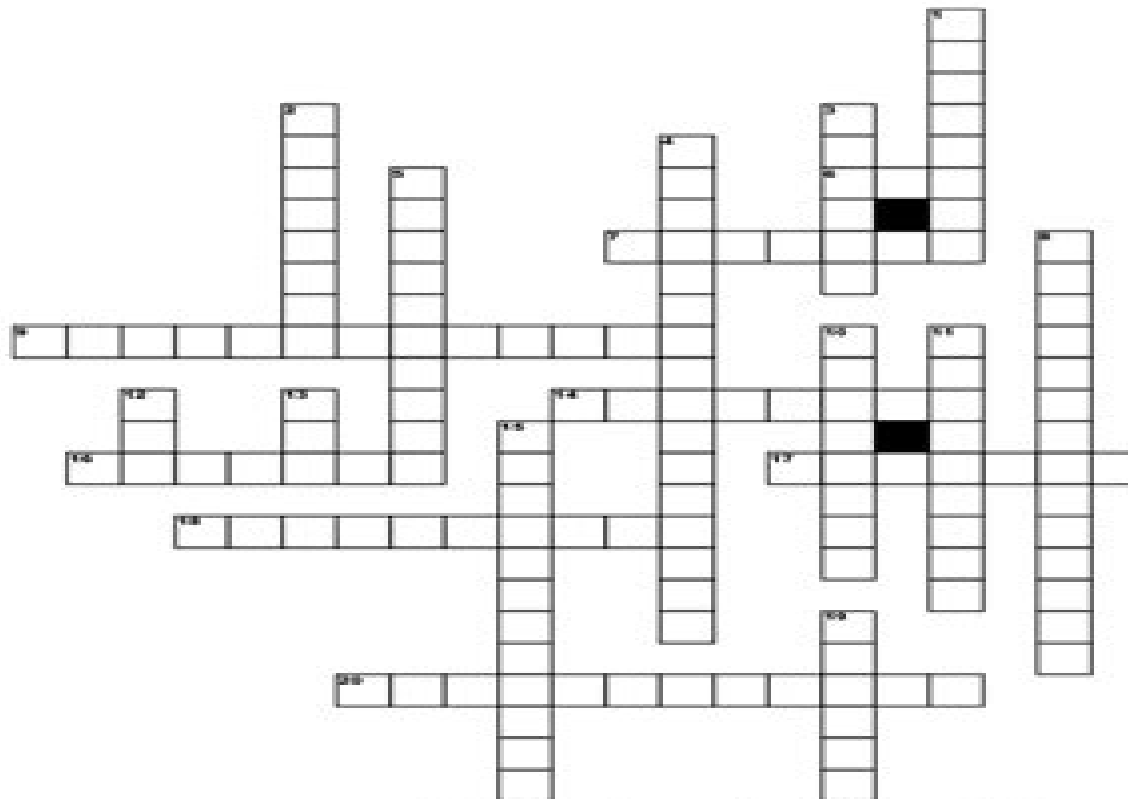


Structure of the Atom



Across

6. Mass Number - Atomic Number = Number of Neutrons
 7. The horizontal rows in the periodic table
 9. Arrangement of elements according to atomic number on a table
 14. A single atom or several atoms bound together electro magnetically bonded.
 16. Center of atom
 17. Positive particles around the nucleus
 18. The sum of the number of neutrons and protons

20. 1. All matter consists of minuscule particles called atoms.
 2. All atoms of a given element are identical to each other. 3. All atoms of a given element are different than those of other elements. 4. Atoms of one element combine with other elements to create compounds. They always combine in equal amounts.

Down

1. Pure substances that can't be broken down or changed
 2. 2 or more elements chemically bonded
 3. Abbreviation for each element
 4. The electrons on the very outer ring of the atom

5. The property of an atom that causes it to have weight
 8. Located outside the nucleus with electrons
 10. Neutral charged particles around the nucleus
 11. Negative particles outside the nucleus
 12. Atomic Mass Unit
 13. Atomic Number = Protons and Electrons
 15. Quantity of protons and electrons in the nucleus of an atom
 19. The vertical rows in the periodic table

Section 3 Reinforcement Periodic Table Word Search Answers

Robert R. Redfield



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