



Reinforcement

Chemical Changes

Read the passage and answer the following questions:



1. What is the chemical formula of hydrogen peroxide?

It is represented as the decomposition of hydrogen peroxide. In this chemical reaction, the reactants are hydrogen peroxide (H_2O_2) and the products are water (H_2O) and oxygen (O_2). The reaction is exothermic, meaning it releases heat and energy in the form of a chemical bond.

2. What is the chemical formula of water? Is it a liquid, solid, or gas? Is it a molecule or a compound? Is it a pure substance or a mixture?

3. What is the chemical formula of oxygen? Is it a liquid, solid, or gas? Is it a molecule or a compound? Is it a pure substance or a mixture?

Hydrogen peroxide is a colorless, odorless liquid. It is a chemical compound made of two hydrogen atoms and two oxygen atoms. It is a pure substance and is not a mixture. It is a liquid at room temperature and pressure. It is a chemical compound and is not a mixture. It is a pure substance and is not a mixture.

4. According to the law of conservation of mass, if hydrogen peroxide decomposes into water and oxygen, what is the mass of the products compared to the mass of the reactants?

It is equal.



5. What is the chemical formula of water? Is it a liquid, solid, or gas? Is it a molecule or a compound? Is it a pure substance or a mixture?

6. What is the chemical formula of oxygen? Is it a liquid, solid, or gas? Is it a molecule or a compound? Is it a pure substance or a mixture?

7. What is the chemical formula of hydrogen peroxide? Is it a liquid, solid, or gas? Is it a molecule or a compound? Is it a pure substance or a mixture?

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8. What is the chemical formula of water? Is it a liquid, solid, or gas? Is it a molecule or a compound? Is it a pure substance or a mixture?

9. What is the chemical formula of oxygen? Is it a liquid, solid, or gas? Is it a molecule or a compound? Is it a pure substance or a mixture?

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10. According to the law of conservation of mass, if hydrogen peroxide decomposes into water and oxygen, what is the mass of the products compared to the mass of the reactants?

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Patrick Vollmar



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