# Questions 13-15 refer to the information below.

- A 0.50 g sample of Mg(s) was placed in a solution of HCl(aq), where it reacted completely.
- 13. Which of the following equations best represents the reacting species in the reaction described above?
  - (A)  $Mg(s) + 2 H^{+}(aq) \rightarrow Mg^{2+}(aq) + H_{2}(g)$
  - (B)  $Mg(s) + HCl(aq) \rightarrow MgCl_2(aq) + H_2(g)$
  - (C)  $Mg(s) + 2 HCl(aq) \rightarrow MgCl_{s}(s) + 2 H^{+}(aq)$
  - (D)  $Mg(s) + 2 HCl(aq) \rightarrow Mg^{2+}(aq) + Cl_2(g) + H_2(g)$
- 14. In another experiment, a small piece of Mg(s) is weighed, then placed in a flask containing excess 1 M HCl(aq). The student wants to determine number of moles of gas produced. Which of the following is the best way to conduct the experiment for accurate data collection?
  - (A) Conducting the experiment at different temperatures to see which generates the most gas
  - (B) Completing the entire reaction in a large Erlenmeyer flask of known volume to measure the volume of the gas collected
  - (C) Collecting the gas in a eudiometer tube and measuring the volume of the gas collected
  - (D) Conducting the reaction in a graduated cylinder and measuring the volume of the gas collected
- 15. In a third experiment, 0.10 g samples of Mg(s) are placed in excess HCl(aq) of various concentrations: 0.050 M, 0.10 M, 0.25 M, and 0.50 M. The reactions are run in successive order from 0.050 M to 0.50 M, and the time required for each reaction to go to completion is recorded. As the concentration of HCl(aq) increases from 0.050 M to 0.50 M, which of the following is the expected result?
  - (A) The reaction time increases, and the rate of the reaction decreases.
  - (B) The reaction time decreases, and the rate of the reaction increases.
  - (C) Both the reaction time and the rate of the reaction increase.
  - (D) Both the reaction time and the rate of the reaction decrease.

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