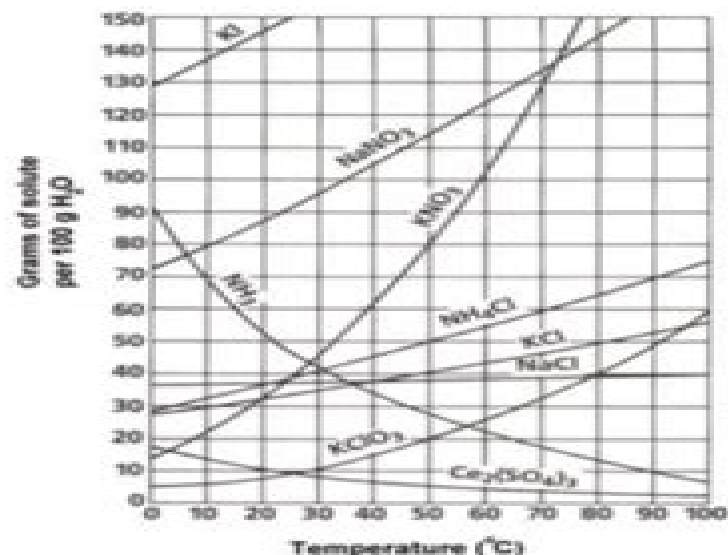


Chemistry I-H
Solubility Curves and Solutions Review Sheet
Solution Set

1. Define solubility: **The maximum amount of solute that dissolves in a fixed quantity of a particular solvent at a specified temperature.**

Use the solubility curve below to answer questions #2 - #15. Be sure to note the units on the axes of the graph.



2. In general, how does temperature affect solubility?
The higher the temp. the more substance dissolves
3. **KClO₃** Which compound is least soluble at 10°C?
4. **50 g** How many grams of KCl can be dissolved in 100g of water at 80°C?
5. **40 g** How many grams of NaCl can be dissolved in 100g of water at 90°C?
6. **about 60 g** At 40°C, how much KNO₃ can be dissolved in 100g of water?
7. **NaCl** Which compound shows the least amount of change in solubility from 0°C-100°C?
8. **unsaturated** At 30°C, 90g of NaNO₃ is dissolved in 100g of water. Is this solution saturated or unsaturated?
9. **D** At 60 C, 72 g of NH₄Cl are dissolved in 100g of water. This solution is considered to be:
A) saturated; B) unsaturated; C) supersaturated; D) saturated with some left undissolved
10. **30 g** A saturated solution of KClO₃ is formed from one hundred grams of water. If the saturated solution is cooled from 90°C to 50°C, how many grams of precipitate are formed?
11. **20 g** A saturated solution of NH₄Cl is formed from one hundred grams of water. If the saturated solution is cooled from 80°C to 40°C, how many grams of precipitate are formed?
12. **E** Which compounds show a *decrease* in solubility from 0°C-100°C?
A) potassium chlorate B) cerium sulfate C) sodium nitrate D) ammonia E) Both B and D
13. **C** Which compound is the most soluble at 10°C?
A) Ce₂(SO₄)₃ B) NaCl C) KCl D) KClO₃ E) NH₃
14. **C** A chemist adds 120 grams of KCl into a test tube which contains 100 grams of water. The temperature of the water is 80°C. How much of the solute will be left undissolved at the bottom of the test tube?
A) 50 grams B) 60 grams C) 70 grams D) 80 grams E) 56 grams

Science Chemistry Unit Solutions Review

Ying-Ying Zheng



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