

GENERAL INFORMATION

GENERAL SPECIFICATIONS

MODEL: *RANGER XP 4X4*

Category	Dimension / Capacity
Length	113 in. / 287 cm
Width	60 in. / 152.4 cm
Height	75 in. / 190.5 cm
Wheel Base	76 in. / 193 cm
Ground Clearance	11.5 in. / (29 cm)
Turning Radius	132 in. / 335 cm
Dry Weight	1185 lbs. / 537 kg
Gross Vehicle Weight	2750 lbs. / 1247 kg
Cargo Box Capacity	1000 lbs. / 454 kg
Cargo Box Dimensions	58 x 42 x 10 in. (147 x 106.7 x 25.4 cm)
Vehicle Payload	1500 lbs. / 681 kg (Includes driver and two passengers - 500 lbs. / 227kg)
Hitch Towing Capacity	1500 lbs. / 681 kg
Hitch Tongue Capacity	150 lbs. / 68 kg



MODEL: *RANGER XP 6X6*

Category	Dimension / Capacity
Length	120 in. / 305 cm
Width	60 in. / 152.4 cm
Height	75 in. / 190.5 cm
Wheel Base	90 in. / 229 cm
Ground Clearance	7.2 in. / (18.25 cm)
Turning Radius	180 in. / 457 cm
Dry Weight	1410 lbs. / 640 kg
Gross Vehicle Weight	2900 lbs. / 1315 kg
Cargo Box Capacity	1250 lbs. / 567 kg
Cargo Box Dimensions	58 x 48 x 10 in. (147 x 122 x 25.4 cm)
Vehicle Payload	1750 lbs. / 794 kg (Includes driver and two passengers - 500 lbs. / 227kg)
Hitch Towing Capacity	1750 lbs. / 794 kg
Hitch Tongue Capacity	150 lbs. / 68 kg



Polaris Ranger 700 6x6 Factory Service Repair Manual

N Colangelo



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web objective to apply stoichiometric principles to calculate the appropriate chemical equation for the reaction between iron metal and a solution of copper i sulfate introduction in the form of finely split powder copper metal is produced as a byproduct of the process as iron ions the iron metal dissolves

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web purpose to enhance the understanding of stoichiometry a reaction between iron and copper ii sulfate solution will be conducted this will help you to differentiate limiting and excess reactant in a chemical reaction finally the theoretical and percent yield of this reaction will be calculated

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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8.00 grams of copper ii sulfate onto a piece of filter paper taking the filter paper's mass into account add these to the beaker 3 measure 50.0 ml of water in the cylinder and add it to the beaker 4

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web in this experiment we will use stoichiometric principles to deduce the appropriate equation for the reaction between metallic iron and a solution of copper ii sulfate this reaction produces metallic copper which is seen precipitating as a finely divided red powder

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web in this lab we will use experimental evidence and the principles of stoichiometry to determine which ion is formed in a single replacement reaction between iron and copper ii sulfate pre lab assignment in preparation for the lab activity we will investigate another single replacement reaction between copper and silver nitrate

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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction

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web this experiment is focused on a reaction between iron metal and a solution of copper ii sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper ii sulfate ions the copper ii is then reduced and gains electrons

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web stoichiometry can be defined as the relationship between the relative quantities of substances taking part in a reaction or forming a compound stoichiometric principles can be used to deduce the appropriate chemical equation for a reaction in

our experiment we used solid iron in the form of powder was added to a solution of copper ii sulfate

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web under the finalization of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7 0002g of copper ii sulfate and 2 0101g iron metal which produced 2 3037g of copper product with a percent yield of 100 72 introduction

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